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Development of Highly Selective Ligands for Separations of Actinides from Lanthanides in the Nuclear Fuel Cycle

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Dedication – We dedicate this paper to the memory of Dr. Charles Madic.

Abstract: This account summarises recent work by ourselves and others on the development of ligands for the separation of actinides from lanthanides contained in nuclear waste streams in the context of a future European strategy for nuclear waste management. The current status of actinide/lanthanide separations worldwide is briefly discussed and the synthesis, development and testing of different classes of soft N- and S-donor heterocyclic ligands in Europe over the last 20 years is presented. This work led to the current benchmark ligand that displays many of the desirable qualities for industrial use. The improvement of radiolytic stability through ligand design is also discussed.

1 Introduction and Background

2 Dithiophosphinic Acids

3 TERPY and Related Ligands

4 TPTZ and Related Ligands

5 BODO and Related Ligands

6 Podant Ligands

7 BTP and Related Ligands

8 BTBP and Related Ligands

9 Improving Radiolytic Stability through Ligand Design

10 Conclusions and Outlook

Key words: americium, heterocycles, ligands, actinides, nitrogen, solvent extraction, europium, separation, lanthanides

1 Introduction and Background

The generation of electricity through nuclear power is becoming increasingly important in order to meet future energy demands while minimizing the production of greenhouse gases.¹ As a result, the inventory of nuclear waste worldwide and thus its environmental burden will continue to increase.² However, an effective strategy for the improved safe long-term management and disposal of the waste remains to be further studied.

The spent nuclear fuel produced by a typical light water reactor is composed of uranium and short-lived fission products such as lanthanides (>98.5 wt %) which do not pose a long term hazard, together with plutonium and the minor actinides Am, Cm and Np (approx. 1 wt %) which are highly radiotoxic. The PUREX (Plutonium and URanium EXtraction) process using tributyl phosphate **1** (Figure 1) is currently used to separate the plutonium from the remaining waste enabling its recycling and reuse as MOX fuel in nuclear reactors.³ The remaining high level liquid waste (PUREX raffinate) still contains the minor actinides. The long-term radioactivity ($T_{1/2}$ =

10^3 – 10^4 years) of these elements presents a serious environmental and economic problem for the safe management and storage of the waste and necessitates that it be isolated from the biosphere for several thousands of years.

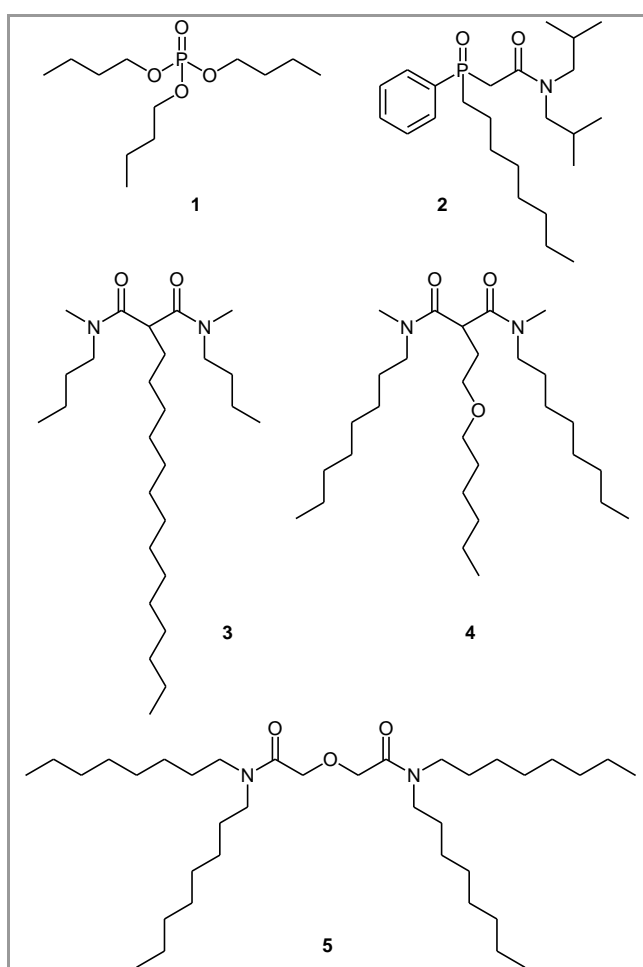


Figure 1 Structures of ligands used in the PUREX, TRUEX and DIAMEX processes

One possible strategy to reduce the radiotoxicity of the waste involves neutron bombardment (transmutation) of the minor actinides into shorter-lived ($T_{1/2}$ = 10^1 years) or stable elements.⁴ However, the lanthanides have high neutron capture cross sections and absorb the neutrons in preference to the transmutable actinides. It is therefore necessary to first separate (partition) the actinides from the bulk of the

lanthanides prior to their transmutation. This strategy, known as Partitioning and Transmutation,⁵ is a promising alternative to long term deep geological disposal of the high level waste.

In recent decades, there has been intensive research worldwide on the development of solvent extraction processes for the partitioning of actinides from lanthanides as part of a future strategy for the reprocessing of spent nuclear fuel.⁶ Within Europe, this research has led to the development of two proposed sequential processes. In the first process, both actinides and lanthanides are co-extracted from PUREX raffinate by a non-selective, diamide-based ligand. This is known as the DIAMEX (DIAMide EXtraction) process.⁷ In the second process, known as the SANEX (Selective ActiNide EXtraction) process, the minor actinides will be selectively extracted and separated from the lanthanides. Currently, the TRUEX (TRAnsuranium EXtraction) process, based on octyl(phenyl)-*N,N*-diisobutylcarbamoylmethylphosphine oxide (CMPO **2**, [Figure 1](#)) as the principal extractant, is used in the United States⁸ to recover the actinides and lanthanides from high level waste solutions but no process is currently used worldwide for the actinide/lanthanide group separation.

Within the framework of several European Union funded research projects, a number of substituted diamide ligands have been developed for the DIAMEX process.⁹ Among these, *N,N'*-dimethyl-*N,N'*-dibutyltetradecylmalonamide (DMDBDTMA) **3**¹⁰ and *N,N'*-dimethyl-*N,N'*-dioctylhexoxyethylmalonamide (DMDOHEMA) **4**¹¹ have shown the most promising results in extraction tests on genuine fuel solutions. More recently, *N,N,N',N'*-tetraoctyldiglycolamide (TODGA) **5** ([Figure 1](#)) showed superior performance than **3** or **4** when tested on PUREX raffinate and this extractant is currently the preferred ligand for the DIAMEX process.¹²

For the SANEX process, the separation of actinides from lanthanides is much more challenging due to the chemically similar behavior of the two groups of elements (identical oxidation states, similar ionic radii due to lanthanide/actinide contractions).¹³ However, there is a more covalent contribution to metal-ligand bonding with the actinides than with the lanthanides. Consequently, ligands with softer donor atoms (eg: N and S) than favour complexation with the 5f orbitals of the actinides have been the focus of research so as to exploit this difference between the coordination chemistries of the actinides and lanthanides. In the United States, the TALSPEAK process has been proposed to carry out this separation.¹⁴

Ligands suitable for the separation of actinides from lanthanides in a SANEX process must obey several important criteria.¹⁵ They should have a sufficiently high solubility in the organic diluent used for the extraction. The extracted species must also have good

solubility in the organic phase to prevent third phase formation or precipitation. The ligand must show a high enough selectivity for the actinides so that the separation can be carried out in a relatively small number of extraction stages. It is also desirable that the affinity for the actinides is not so high that the back-extraction (stripping) of the metal from the extracted complexes and the subsequent recycling of the organic phase cannot take place. The ligand must show acceptable resistance towards both acidic hydrolysis and radiolysis, and any degradation products that form must not interfere with the extraction. The ligand should be able to extract from highly acidic (up to 4 M HNO₃) solutions. It is also desirable that the ligand is composed of only C, H, O and N (CHON principle) so that the ligand or its degradation products can be completely incinerated at the end of its useful life, thereby avoiding secondary waste generation. Finally, the ligand must be as cheap and easy to synthesize as possible so that large-scale commercial production is feasible.

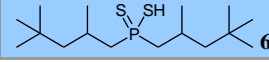
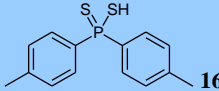
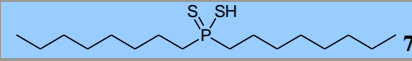
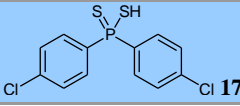
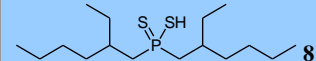
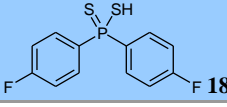
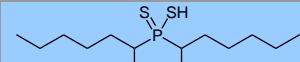

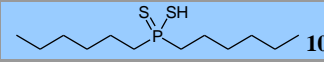

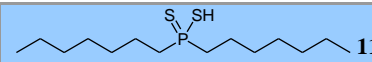
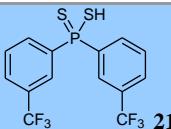
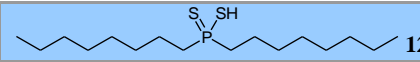
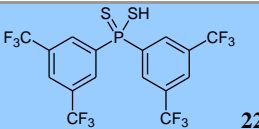
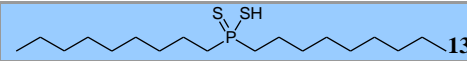
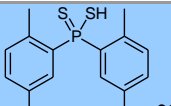
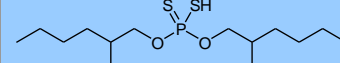
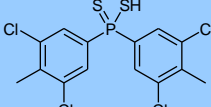
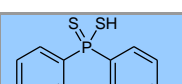
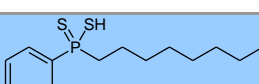
In solvent extraction experiments, ²⁴¹Am(III) and ¹⁵²Eu(III) radioisotopes are typically used as representatives for the actinides and lanthanides, respectively. The effectiveness of a ligand in extracting a metal M is quantified by the distribution ratio (D_M), which is the equilibrium ratio between the metal concentration in the organic and in the aqueous phase. The separation factor for Am over Eu ($SF_{Am/Eu}$) is the ratio of the D values for Am and Eu and is thus a measure of the selectivity of a ligand for the actinides over the lanthanides. Thus the optimum ligand will have $D_{Am} > 1$ and $SF_{Am/Eu}$ as large as possible.

A large number of potential SANEX ligands has been developed in recent decades.¹⁶ In this account, we review the N- and S-donor ligands for the SANEX process that have been developed within the framework of several European Union funded research programmes that show the most promising properties for actinide/lanthanide separations. The harder O-donor ligands, which are generally non-selective, are thus out of the scope of this review.

2 Dithiophosphinic Acids

The bidentate, S-donor dithiophosphinic acids were one of the first classes of reagent to show very high selectivities for An(III) over Ln(III). Dithiophosphinic acids studied for their ability to separate An from Ln are shown in [Table 1](#). It was shown in early studies by Zhu that Cyanex 301, a commercial reagent composed mainly (~ 80 %) of bis(2,4,4-trimethylpentyl)dithiophosphinic acid **6**, efficiently separated Am(III) from Ln(III) with very high selectivities. The selectivity of the extraction increased when the reagent was fully purified ($SF_{Am/Eu} > 4000$, average $SF_{Am/Ln} > 2300$).¹⁷ Both the

Table 1 Dithiophosphinic acids studied for their ability to extract and separate An(III) from Ln(III)

Ligand Structure	References	Ligand Structure	References
 6	17, 18, 19, 20, 21, 22, 23, 24, 25, 35	 16	26b, 27, 33
 7	33	 17	26, 27, 28, 29, 30, 31, 33, 34
 8	33	 18	26a, 27
 9	33	 19	32
 10	33	 20	32b
 11	33	 21	32b
 12	33	 22	32a
 13	33	 23	33
 14	33	 24	33
 15	26a, 26b, 27, 33, 35	 25	33

extraction efficiency and selectivity were enhanced in formed mixed metal complexes with **6**.¹⁸ In synergistic combination with the softer N-donor ligands 2,2'-bipyridine and 1,10-phenanthroline, the selectivity was further enhanced resulting in one of

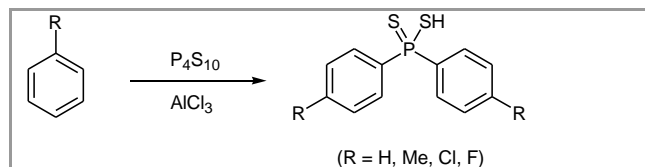
the presence of several O-donor synergists which the highest reported separation factors to date ($SF_{Am/Eu} > 40000$).¹⁹

Promising results were also observed in the extraction

of An(III) and Ln(III) from acetic acid/ammonium acetate buffered solutions that would result from a preceding TRUEX separation step.²⁰ In addition, the successful separation of An(III) from Ln(III) by purified bis(2,4,4-trimethylpentyl)dithiophosphinic acid **6** immobilized on silica resins,²¹ or enclosed in biopolymer microcapsules²² or in supported liquid membranes,²³ has been demonstrated. In a laboratory-scale countercurrent extraction test, > 99.998 % of Am(III) was extracted and separated from macro-amounts of Ln(III) using purified **6** and tributyl phosphate.²⁴ Unfortunately, **6** can only extract from nitric acid solutions of relatively low acidity (pH > 2.5) due to non-dissociation/oxidative degradation of the extractant at low pH. Furthermore, **6** was found to decompose to its corresponding monothiophosphinic acid and phosphinic acids on exposure to 700 Kgy of γ -radiation, with deleterious effects on the extraction performance and selectivity.²⁵

To address these shortcomings, a series of more acidic aromatic dithiophosphinic acids was developed by Modolo. Compounds **15–18** were synthesized by treatment of phosphorus pentasulfide with aluminum chloride in the presence of an aromatic hydrocarbon (Scheme 1). The ligands efficiently extracted Am(III) and Eu(III) into toluene from 0.1–1 M HNO₃ in synergistic combination with tributyl phosphate or a trialkylphosphine oxide, with the best results being observed for bis(chlorophenyl)dithiophosphinic acid **17**.²⁶ No extraction occurred by the ligands alone in the absence of a synergist. The dependence of the distribution ratios for An(III) on extractant concentration indicate that the dithiophosphinic acids extract An(III) as 1:3 complexes. The aromatic dithiophosphinic acids were also more resistant to radiolysis than Cyanex 301 **6**. Only minor (< 5 %) decomposition was observed after exposure to 1000 Kgy of γ -radiation, whereas > 80 % decomposition of Cyanex 301 **6** occurred under the same conditions.²⁷ However, complete oxidative degradation of the ligands occurred in contact with strongly acidic (> 2 M) nitric acid solution over 100 days.²⁸

The synergistic mixture of bis(chlorophenyl)dithiophosphinic acid **17** and tris(2-ethylhexyl)phosphate extracted Am(III) and Cm(III) from the lanthanides with very high separation factors ($SF_{Am/Ln} > 3000$). More surprisingly, an Am(III)/Cm(III) separation factor > 8 was observed, suggesting that the system could be used for the challenging separation of Am(III) from Cm(III).²⁹ In a subsequent laboratory-scale separation process using centrifugal contactors, this synergistic combination selectively extracted 99.8 % Am(III) from a mixture of Am(III), Cm(III), Cf(III) and Eu(III), with only 0.47 % Cm(III) being extracted into the organic phase.³⁰



Scheme 1 Synthesis of aromatic dithiophosphinic acids

It has been shown that a donor-acceptor mechanism of extraction operates to promote the extraction of Am(III) by synergistic mixtures of dithiophosphinic acids and neutral O-donor ligands.³¹ Strong O-donor ligands such as trialkylphosphine oxides tend to strengthen M–S bonds and increase D_{Am} , whereas weaker O-donor ligands such as triaryl phosphates tend to weaken M–S bonds and decrease D_{Am} . For those synergists giving high Am/Eu separation factors (trialkyl phosphates), an antagonistic effect was observed for Eu(III) complex formation but not for Am(III) complex formation.

Recently, bis(trifluoromethylphenyl)dithiophosphinic acids **19**, **20** and **21** were synthesized regioselectively by a multi-step route and evaluated in Am(III) and Eu(III) solvent extraction experiments to determine the effect of the location of the trifluoromethyl-group on the extraction properties.³² The *ortho*-substituted ligands **19** and **20** showed unprecedented selectivity for Am(III) over Eu(III) ($SF_{Am/Eu} > 100,000$) but the *meta*-disubstituted ligand **21** was far less selective ($SF_{Am/Eu} \sim 20$), demonstrating that the location of the substituent played a key role in the extraction selectivity. In a related study on several alkyl and aryl-substituted dithiophosphinic acids, the presence of electron-withdrawing groups on the aromatic rings in the aryl-substituted ligands increased the selectivity, but decreased the extraction ability.³³ For the alkyl-substituted ligands, branched alkyl chains increased the selectivity.

The origin of the high selectivity of synergistic mixtures of bis(chlorophenyl)dithiophosphinic acid **17** and neutral O-donor ligands for An(III) over Ln(III) was established by EXAFS and time-resolved laser fluorescence spectroscopy (TRLFS).³⁴ The structures of the inner coordination spheres of the Cm(III) and Eu(III) complexes were different. While Cm(III) primarily bound to the S-atoms of the ligand in a bidentate fashion, Eu(III) preferentially bound the O-atoms of the synergist, suggesting that a high number of covalent M–S bonds is responsible for the high selectivity. The primary role of the synergist appears to be co-complexation of the metal ions to form more hydrophobic complexes. Molecular dynamics simulations show that the ligands and their Eu(III) complexes are surface active, but synergists like tributyl phosphate also promote the extraction to the organic phase by aggregating at the interface where they lower its surface tension and promote the transfer of the complexes into the organic phase.³⁵

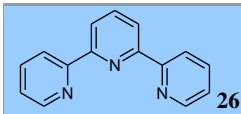
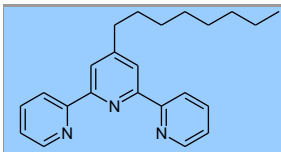
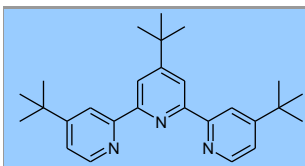
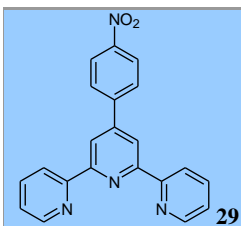
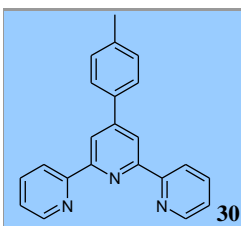
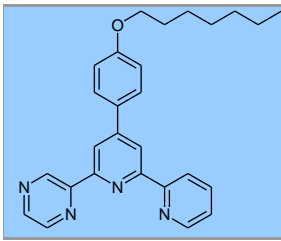
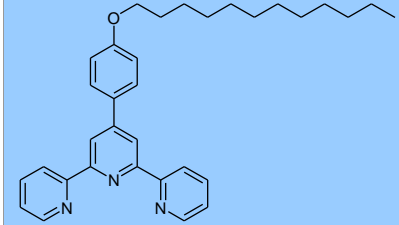
3 TERPY and Related Ligands

The tridentate aromatic nitrogen donor ligand 2,2':6',2''-terpyridine **26** (TERPY) was one of the first N-donor ligands investigated for its ability to selectively extract Am(III) from Eu(III). A list of TERPY and its derivatives tested for their ability to separate Am(III) from Eu(III) is shown in Table 2. In synergistic combination with 2-bromodecanoic acid (which acts as a lipophilic anion source), TERPY selectively extracted Am(III) from weakly acidic aqueous solutions up to 0.1 M HNO₃ into THP (hydrogenated tetrapropene) with SF_{Am/Eu} of 7.2.^{36,37} However, the *D* value for Am(III) progressively decreased at higher acidities ([HNO₃] > 0.1) due to protonation and subsequent dissolution of the ligand in the aqueous phase. No significant extraction of Am(III) or Eu(III) was observed by TERPY alone.

The more hydrophobic derivatives 4'-octyl-2,2':6',2''-terpyridine **27** and 4,4',4''-tri-*tert*-butyl-2,2':6',2''-terpyridine **28** (Table 2) were then investigated. However, at comparable acidities the *D* values for Am(III) were substantially lower than for TERPY **26**.³⁶ Although alkyl substitution was expected to increase the *D* values by improving the solubility of the extracted complex in the organic phase, it appeared to increase the basicity of the ligand so that protonation of the donor atoms competed effectively with metal ligation. Similar results were observed with phenyl-substituted derivatives **29**, **30** and **32**, and pyrazine derivative **31**.³⁸ All of the substituted ligands became protonated at higher acid concentrations and precipitation of the protonated ligands was observed in some cases.

In an effort to better understand the types of species which may be involved in the extraction, a series of complexes formed by TERPY **26** and the lanthanides were synthesized and structurally characterized.^{38,39} In most cases, neutral 1:1 species [M(L)(NO₃)₃] (L = TERPY) were observed although with Nd, Sm, Tb, Dy and Ho, complex ion-pairs composed of [M(L)₂(NO₃)₂]⁺ cations and [M(L)(NO₃)₄]⁻ anions were observed. However, under anhydrous conditions 1:2 complexes of formula [MI₂(TERPY)₂]I were formed between TERPY and MI₃ (M = Ce, Nd, U).⁴⁰ X-ray crystallographic analysis revealed shorter average M–N distances in the U complex than the Ce complex, suggesting a possible π -backbonding interaction and thus greater covalency in the U(III) complex.

Table 2 TERPY and related ligands studied for their ability to extract and separate An(III) from Ln(III)

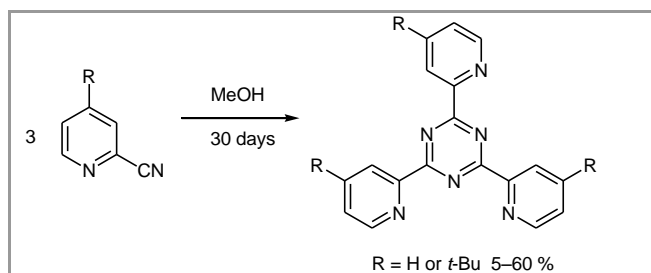
Ligand Structure	References
	36, 37, 38, 39, 40, 41, 46, 73b, 76b, 78, 109b
	36
	36
	38
	38
	38
	38

The structures of the mono- and diprotonated TERPY cations were determined by *Ab initio* quantum

mechanics calculations and by X-ray crystallographic analysis of the ion pairs formed between the diprotonated TERPY cation and nitrato lanthanate ($[M(NO_3)_5]^{2-}$) anions of La and Sm.⁴¹ The isolation of these species in the solid state suggests that TERPY and its derivatives could extract the metals by an ion-pair mechanism as well as by simple solvation, particularly at high acidities.

4 TPTZ and Related Ligands

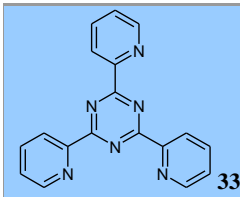
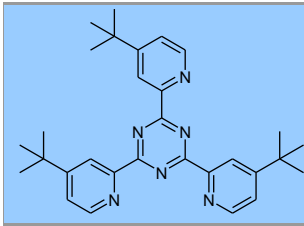
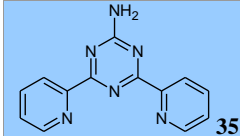
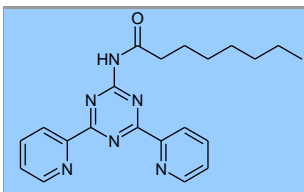
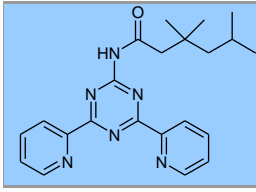
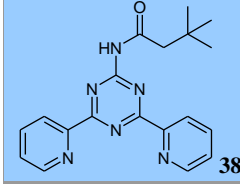
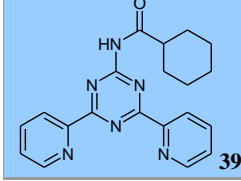
The tridentate N-donor ligand 2,4,6-tri-(2-pyridyl)-1,3,5-triazine (TPTZ) and its derivatives are synthesized by cyclotrimerization of 2-cyanopyridines under acidic, basic or neutral conditions (Scheme 2).⁴² 2,4,6-tri-(2-pyridyl)-1,3,5-triazine and related ligands that were tested for their ability to separate Am(III) from Eu(III) are shown in Table 3. It was reasoned that replacing the central pyridine ring of TERPY with a 1,3,5-triazine ring would reduce the basicity of the ligand and therefore lower the solubility of the protonated ligand in the aqueous phase at low pH.



Scheme 2 Synthesis of 2,4,6-tri-(2-pyridyl)-1,3,5-triazine ligands

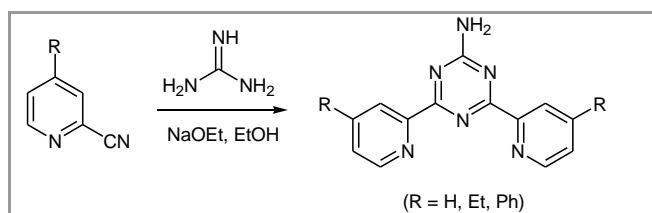
In solvent extraction experiments, TPTZ **33** performed slightly better than TERPY **26** and higher *D* values were observed for Am(III) when used in combination with 2-bromodecanoic acid in TPH. TPTZ was also the first N-donor ligand to show $SF_{Am/Eu} > 10$. In contrast to the more lipophilic derivatives of TERPY, the extraction was further improved in the more lipophilic TPTZ derivative **34** which gave slightly higher *D*_{Am} values than TPTZ and $SF_{Am/Eu} = 11.9$.³⁶ Once again however, the ligands were incapable of extracting the metals from solutions more acidic than 0.1 M HNO₃. X-ray crystallographic structures of the 1:1 complexes formed between Ce(IV) and Sm(III) nitrates and 2,4,6-tri-(4-*tert*-butyl-2-pyridyl)-1,3,5-triazine **34**, and complexes in which the protonated ligand forms salts with a variety of anions have been obtained, illustrating the versatile coordination chemistry of this class of ligand.⁴³ For TPTZ **33**, both 1:1 and 1:2 complexes with the trivalent lanthanides and actinides have been observed under different conditions.⁴⁴

Table 3 TPTZ and related ligands studied for their ability to extract and separate An(III) from Ln(III)

Ligand Structure	References
	36, 37, 42b, 42c, 42d, 43b, 76b, 78, 108
	36, 42c, 43a, 46
	45, 46, 47, 76b, 78
	46
	46
	46
	46

The promising results obtained with TPTZ and its derivatives led us to investigate the related ligand 4-amino-2,6-bis(2-pyridyl)-1,3,5-triazine **35** (ADPTZ) that features the same tridentate coordination cavity as

TPTZ, but is more easily functionalized to obtain more hydrophobic derivatives. This and related ligands are synthesized by treatment of a 2-cyanopyridine with guanidine in ethanolic sodium ethoxide (Scheme 3).^{42a} The coordination chemistry of **35** with the trivalent lanthanides has been investigated.⁴⁵ X-ray crystallographic analysis revealed the formation of five structural types. In each case, the tridentate ligand formed 1:1 complexes and the remainder of the metals inner coordination sphere was occupied by a combination of water molecules and nitrate ions. The structures appeared to be stabilized by intermolecular hydrogen bonding involving the free amino-group.



Scheme 3 Synthesis of 4-amino-2,6-bis(2-pyridyl)-1,3,5-triazine ligands

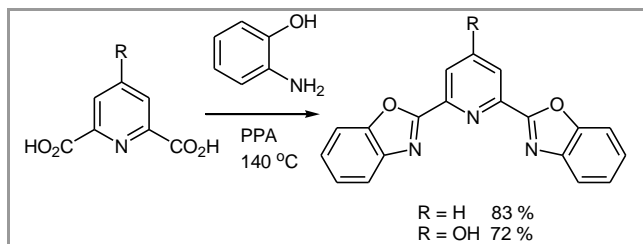
In Am(III)/Eu(III) solvent extraction experiments, **35** and its hydrophobic amido-derivatives **36–39** (Table 3) gave results comparable to TPTZ **33** and **34**.⁴⁶ Although relatively high *D* values for Am(III) were obtained in synergy with 2-bromodecanoic acid (eg: $D_{\text{Am}} = 45$; organic phase = 0.02 M **35** and 1 M 2-bromodecanoic acid in TPH; aqueous phase = 0.03 M HNO₃ with Am(III) and Eu(III) tracers), these values decreased rapidly as the concentration of nitric acid increased. Am(III) distribution ratios less than 1 were observed when [HNO₃] ≥ 0.1 M. The selectivity ($SF_{\text{Am/Eu}}$) ranged between 6.5 and 12.

The amido-derivatives **36–39** showed identical coordination chemistry towards the lanthanides as the parent compound **35**, and crystallographic structures of the 1:1 complexes with Yb(III) have been obtained.⁴⁶ Complexation and microcalorimetry studies in H₂O/MeOH showed that the Am(III) complex of ADPTZ **35** was more stable than the lanthanide complexes by a factor of about 20, and is driven primarily by a more exothermic enthalpy of complex formation.⁴⁷ This supports the view that the preference of the ligand for An(III) is due to a greater degree of covalency in the M–L bonds than with Ln(III).

5 BODO and Related Ligands

In searching for a new class of weakly basic ligand that could also separate An(III) from Ln(III), our attention was drawn to ligands related to 2,6-bis(benzimidazol-2-yl)pyridine whose synthesis takes place in polyphosphoric acid at 200–230 °C.⁴⁸ This stability towards strongly acidic conditions at high

temperatures would be very desirable in a future separation process. In addition, ligands based on the related 2,6-bis(benzoxazol-2-yl)pyridine and 2,6-bis(benzothiazol-2-yl)pyridine would be expected to be less basic, as both benzoxazole and benzothiazole have p*K*_a values lower than benzimidazole. A typical procedure for the synthesis of 2,6-bis(benzoxazol-2-yl)pyridines is shown in Scheme 4. Thus a range of hydrophobic tridentate pyridine ligands containing benzimidazole, benzoxazole and benzothiazole rings (see Table 4) were synthesized and assessed for their ability to extract and separate An(III) from Ln(III).⁴⁹

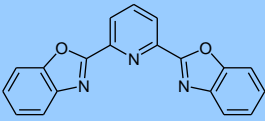
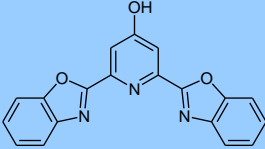
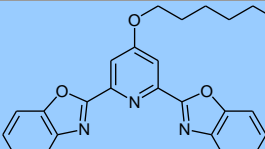
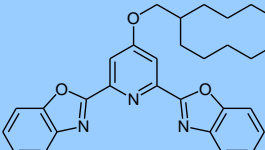
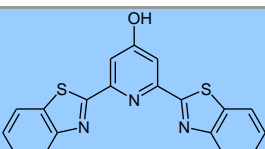
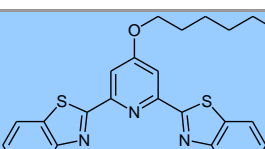
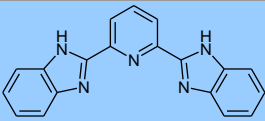
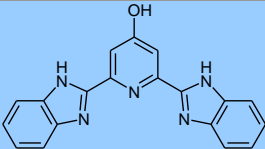


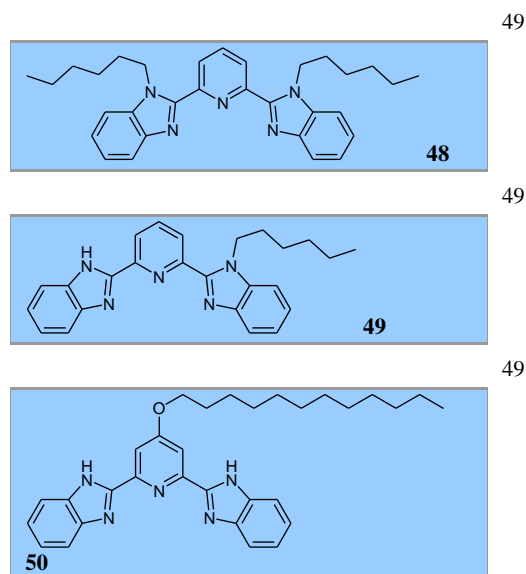
Scheme 4 Synthesis of bis(benzoxazol-2-yl)pyridine ligands

The most promising results were obtained with 2,6-bis(benzoxazol-2-yl)-4-dodecyloxy pyridine (**BODO**) **42** and its branched derivative **43** in combination with 2-bromodecanoic acid as a synergist. Relatively high separation factors were observed ($SF_{\text{Am/Eu}}$ 36–79 for **42**, $SF_{\text{Am/Eu}}$ 43–70 for **43**) and these values actually increased at higher acidities (up to 0.1 M HNO₃). This was the first time that an increase in $SF_{\text{Am/Eu}}$ had been observed with increasing [HNO₃] for an N-donor ligand. The *D* values for Am(III) were greater for the branched ligand **43** than for **42** but once again these values decreased as [HNO₃] increased, and no significant extraction occurred at higher acidities. X-ray crystallographic studies with 2,6-bis(benzoxazol-2-yl)pyridine **40** showed that, as with other terdentate ligands based on TERPY and TPTZ, only 1:1 complex formation occurred with the trivalent lanthanide nitrates.⁴⁹

In further solvent extraction experiments with **BODO**, it was found that at constant acidity, both D_{Am} and $SF_{\text{Am/Eu}}$ decreased as either [NO₃[−]] or [ClO₄[−]] increased.⁵⁰ In addition, the ligand was not extracted into the aqueous phase when contacted with 2 M or 4 M HNO₃ solutions and could be recovered from the organic phase, demonstrating its resistance to protonation. The decrease in D_{Am} and Am/Eu selectivity as the nitrate concentration increases is most likely due to the inability of **BODO** to compete with nitrate ions which must be displaced from the metal by the ligand so that a hydrophobic extractable complex can form.

Table 4 BODO and related ligands studied for their ability to extract and separate An(III) from Ln(III)

Ligand Structure	References
	49
	49
	49, 50
	49
	49
	49, 78
	49
	



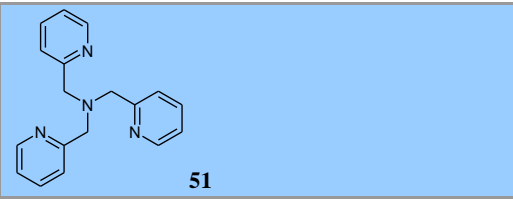
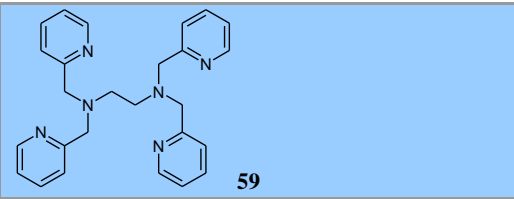
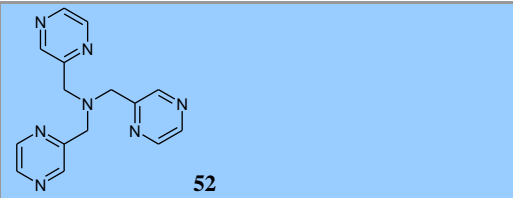
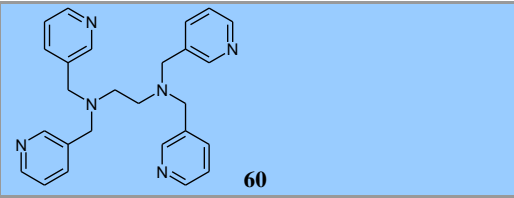
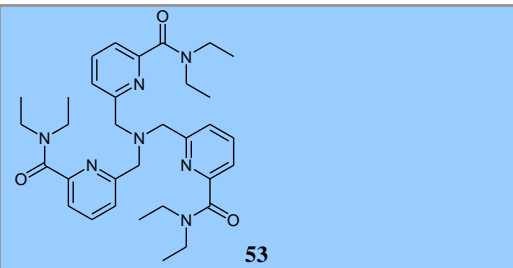
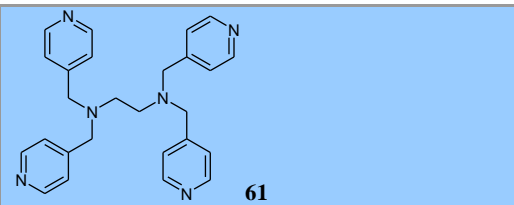
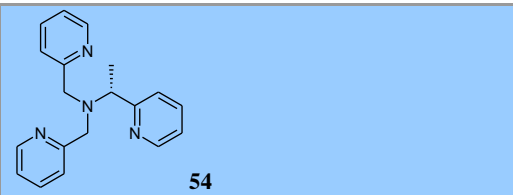
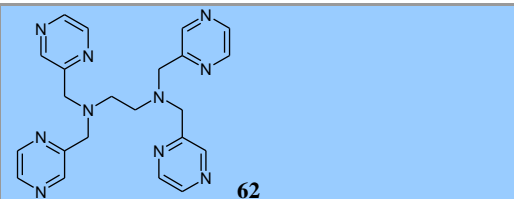
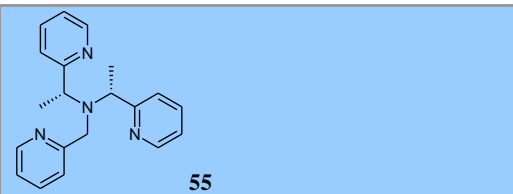
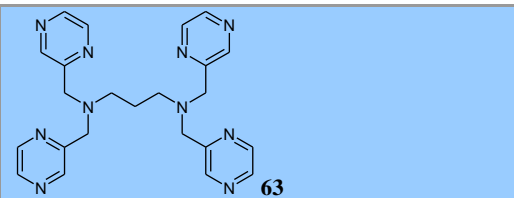
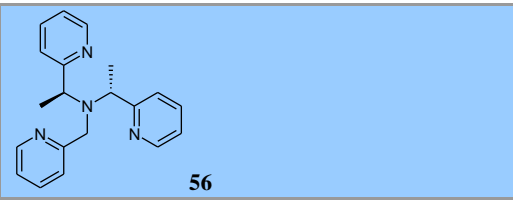
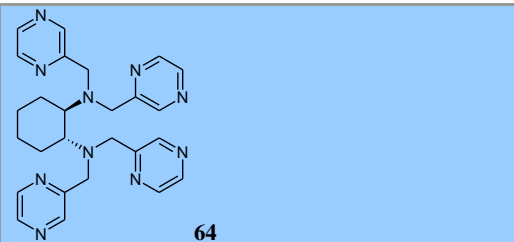
6 Podant Ligands

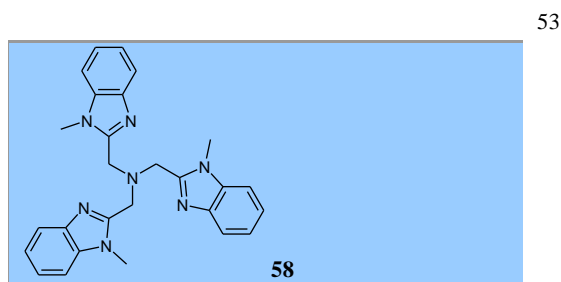
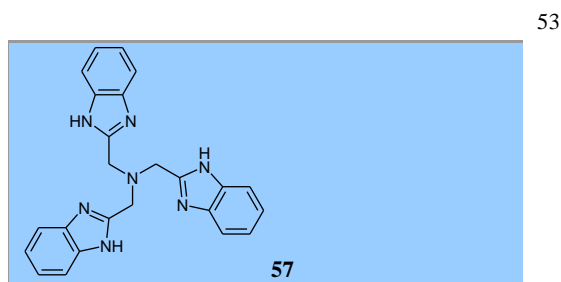
Encapsulating podant ligands are well known in coordination chemistry, and a series of polydentate podant N-donor ligands have been investigated for their ability to separate An(III) from Ln(III). The ligands investigated are presented in Table 5, and are based on the tripodal ligands tris[(2-pyridyl)methyl]amine and tris(2-pyridyl)methanol, and the tetrapodal ligand *N,N,N',N'*-tetrakis[(2-pyridyl)methyl]-ethylenediamine (TPEN).

In solvent extraction experiments, tris[(2-pyridyl)methyl]amine **51** (1 mM in TPH) only weakly extracted Am(III) from nitric acid solutions ($D_{Am} \sim 0.25$ from ca. 0.01 M HNO₃) with low selectivity ($SF_{Am/Eu} < 2$) in the presence of the synergist 2-bromodecanoic acid.⁵¹ The results with the softer, more soluble ligand tris[(2-pyrazinyl)methyl]amine **52** were more promising ($D_{Am} = 4.6$, $SF_{Am/Eu} = 10.8$ for 20 mM **52** in TPH) but no extraction occurred from more acidic solutions ($pH \leq 1$) due to protonation of the ligands. More recently, chiral tris[(2-pyridyl)methyl]amine ligands **54**, **55** and **56** exhibited improved extraction performance at $pH > 4.7$ in nitrobenzene in the presence of 2-bromodecanoic acid.⁵² The best results were obtained with the monomethyl compound (*R*)-**54** ($D_{Am} = 52.3$, $SF_{Am/Eu} \sim 10$) and the dimethyl compound (*R,S*)-**56** ($D_{Am} = 26.8$, $SF_{Am/Eu} > 25$).

A series of 1:1 complexes of tris[(2-pyridyl)methyl]amine **51** and tris[(2-pyrazinyl)methyl]amine **52** with the lanthanides have been isolated and characterized by X-ray crystallography.⁵¹ The structures show the ligand coordinates in a tetradentate capping mode with participation of the central podant N atom. A similar 1:1 complex was formed between tris[(2-

Table 5 Podant ligands studied for their ability to extract and separate An(III) from Ln(III)

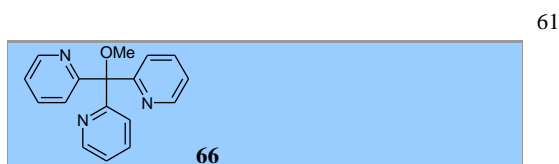
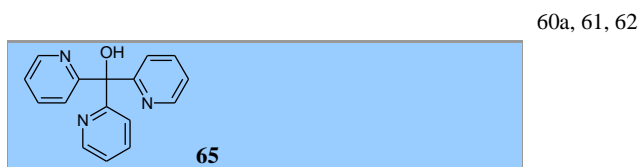
Ligand Structure	References	Ligand Structure	References
	51, 53, 54, 55		57, 58
	51, 56, 59		58
	55		58
	52		59
	52		59
	52		59



pyridyl)methyl]amine **51** and $[\text{UI}_3(\text{THF})_4]$ under anhydrous conditions but the related benzimidazole tripodes **57** and **58** formed 1:2 complexes with U(III).⁵³ Solution NMR studies showed a greater stability of the 1:2 complexes of U(III) and La(III) with these ligands than with tris[(2-pyridyl)methyl]amine **51**, due to the presence of π - π stacking in the benzimidazole complexes. However, successive formation of 1:1 and 1:2 complexes of tris[(2-pyridyl)methyl]amine **51** with the lanthanides was observed on addition of **51** to $\text{Ln}(\text{OTf})_3$ and $[\text{LnI}_3(\text{THF})_4]$ under rigorously anhydrous conditions, and the structures of the 1:2 complexes have been determined.⁵⁴

Analysis of lanthanide-induced proton NMR shifts show that tris[(2-pyridyl)methyl]amine **51** and its triamide derivative **53** have very similar formation constants, although the thermodynamic driving force for complex formation differs in each case. Whereas complexation of tris[(2-pyridyl)methyl]amine **51** is largely enthalpy-driven, complexation with **53** is completely driven by entropy.⁵⁵ The additional ligating sites present in **53** displace more water molecules from the hydrated lanthanide ion during complex formation than **51**. X-ray absorption spectroscopy (EXAFS) on An(III) and Ln(III) complexes of tris[(2-pyrazinyl)methyl]amine **52** showed no obvious structural differences between the two sets of complexes,⁵⁶ although relatively short U(III)-N bond distances were observed, suggesting a higher degree of covalent character.

The tetrapodal, hexadentate ligand *N,N,N',N'*-tetrakis[(2-pyridyl)methyl]-ethylenediamine (TPEN) **59** showed more promising results. In the absence of a synergist, solutions of **59** in nitrobenzene selectively extracted Am(III) over Eu(III) from NH_4NO_3 solutions (pH = 4) as their 1:1 complexes with high efficiency and selectivity ($D_{\text{Am}} \sim 100$, $\text{SF}_{\text{Am/Eu}} \sim 100$ for 10 mM solutions of **59**). The results were broadly

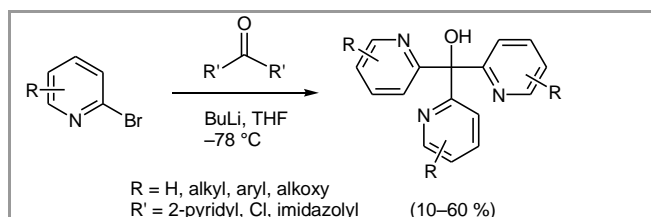


similar in the preferred diluent 1-octanol in the presence of decanoic acid as synergist, although the system did not extract from more acidic (pH < 4) aqueous solutions.⁵⁷ The related TPEN isomers **60** and **61** showed no significant extraction of either Am(III) or Eu(III), indicating the importance of the position of the N-donor atoms in the pyridine rings.⁵⁸

Pyrazine containing analogues of TPEN **59** with different diamino spacers were then investigated. Excellent results were obtained in the extraction of Am(III) and Eu(III) into TPH solution by *N,N,N',N'*-tetrakis[(2-pyrazinyl)methyl]-ethylenediamine **62** in the presence of 2-bromodecanoic acid ($D_{\text{Am}} = 45$, $\text{SF}_{\text{Am/Eu}} = 70$ for 0.1 M **62** from 0.03 M HNO_3). Interestingly, *N,N,N',N'*-tetrakis[(2-pyrazinyl)methyl]-propanediamine **63**, which has a longer spacer and is thus more flexible, shows good extraction ability but no selectivity ($\text{SF}_{\text{Am/Eu}} = 1.5$ –2.3) under similar conditions.⁵⁹ The more rigid ligand **64** bearing a *trans*-1,2-diaminocyclohexane spacer showed no significant extraction of either metal ion. The drastic difference in selectivity between **62** and **63** or **64** was explained by the different preferred conformations of the ligands which was reflected in the X-ray structures of their 1:1 complexes with La(III).

More recently, tris(2-pyridyl)methanol **65** and analogues have been studied as potential An(III)-selective extractants. These ligands are synthesized as shown in Scheme 5 by addition of lithiated 2-bromopyridines to triphosgene, carbonyldiimidazole or bis(2-pyridyl)ketones.⁶⁰ At pH > 4.7, tris(2-pyridyl)methanol **65** selectively extracted Am(III) in synergistic combination with picric acid into nitrobenzene, 1,2-dichloroethane or chloroform with $\text{SF}_{\text{Am/Eu}}$ of 6.7–9.6. The corresponding methyl ether **66** showed similar selectivity but the D values were about three orders of magnitude lower.⁶¹ The ligands extracted Am(III) as 1:3 complexes. Unusual homodinuclear 2:2 lanthanide complexes were obtained on treatment of tris(2-pyridyl)methanol **65**

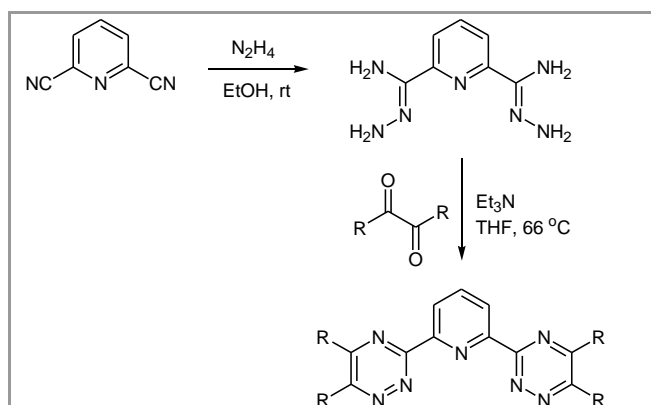
with $\text{La}(\text{OTf})_3 \cdot 3\text{H}_2\text{O}$. Each ligand coordinated in a tetradentate bridging mode with μ -oxo bridges formed by the hydroxyl groups of **65**.⁶²



Scheme 5 Synthesis of tris(2-pyridyl)methanol ligands

7 BTP and Related Ligands

Terdentate ligands based on 2,6-bis(1,2,4-triazin-3-yl)pyridine (BTP) were first synthesized by Case in 1971.⁶³ BTPs are typically synthesized by the condensation reaction of pyridine-2,6-dicarbohydrazonamide with a 1,2-diketone (**Scheme 6**). Related 1,2,4-triazine-substituted pyridines have found use as synthetic precursors to functionalized oligopyridines.⁶⁴ In 1999, Kolarik reported that alkyl-substituted BTPs and related 2,6-bis(1,2,4-triazolyl)pyridines separated Am(III) from Eu(III) with very high selectivity ($\text{SF}_{\text{Am/Eu}} \leq 150$).⁶⁵ The BTP and related ligands that were screened for their ability to selectively extract Am(III) from Eu(III) are summarized in **Table 6**.



Scheme 6 Synthesis of 2,6-bis(1,2,4-triazin-3-yl)pyridine ligands

The extraction by the triazolyl-based ligands depended strongly on the diluent composition used and a synergist (2-bromohexanoic acid) was required for efficient extractions. The related oxadiazolyl-based ligands performed similarly, but with lower extraction selectivity ($\text{SF}_{\text{Am/Eu}} \leq 23.3$).⁶⁶ In contrast, not only were the BTP ligands able to selectively extract Am(III) with very high efficiency ($1 \leq D_{\text{Am}} \leq 61$) in the absence of a lipophilic anion source (ie: $\text{Am}(\text{NO}_3)_3$ was extracted), but they could also extract efficiently from solutions of high acidity ($[\text{HNO}_3] \geq 1 \text{ M}$). This was a key breakthrough because up to this point, other heterocyclic *N*-donor ligands required synergists such

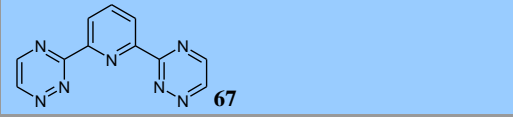
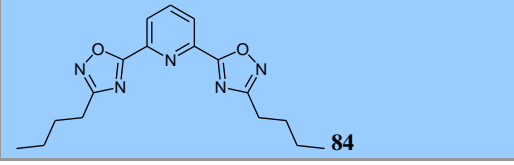
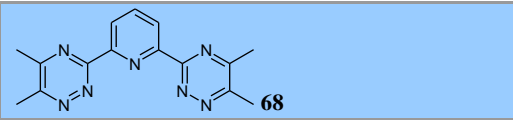
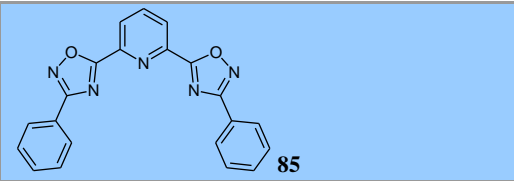
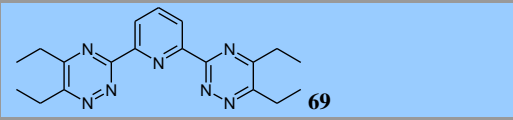
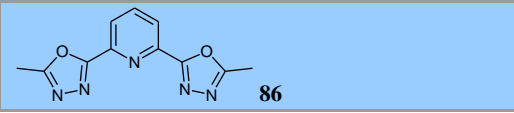
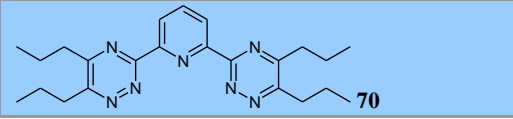
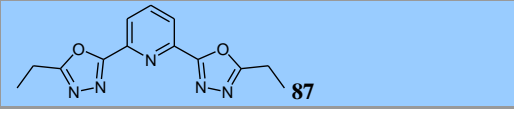
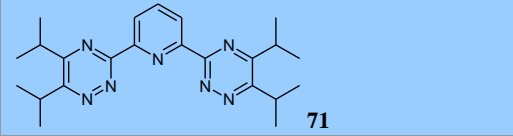
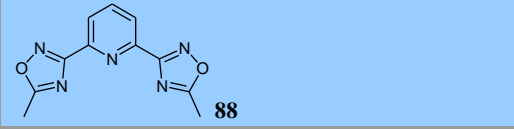

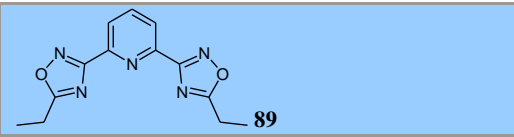
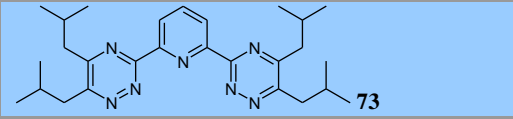
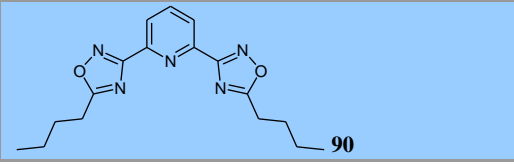
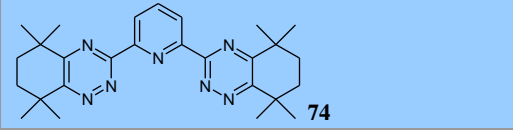
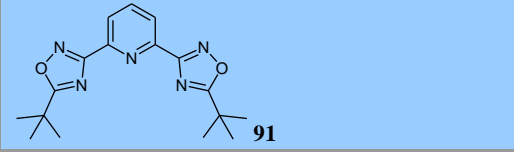
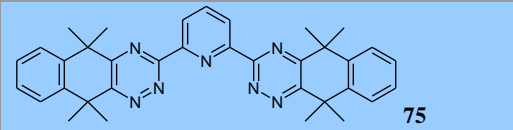
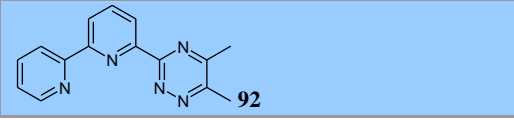
as 2-bromodecanoic acid for efficient extractions, and were unable to extract from highly acidic solutions that would be present in an industrial SANEX process.

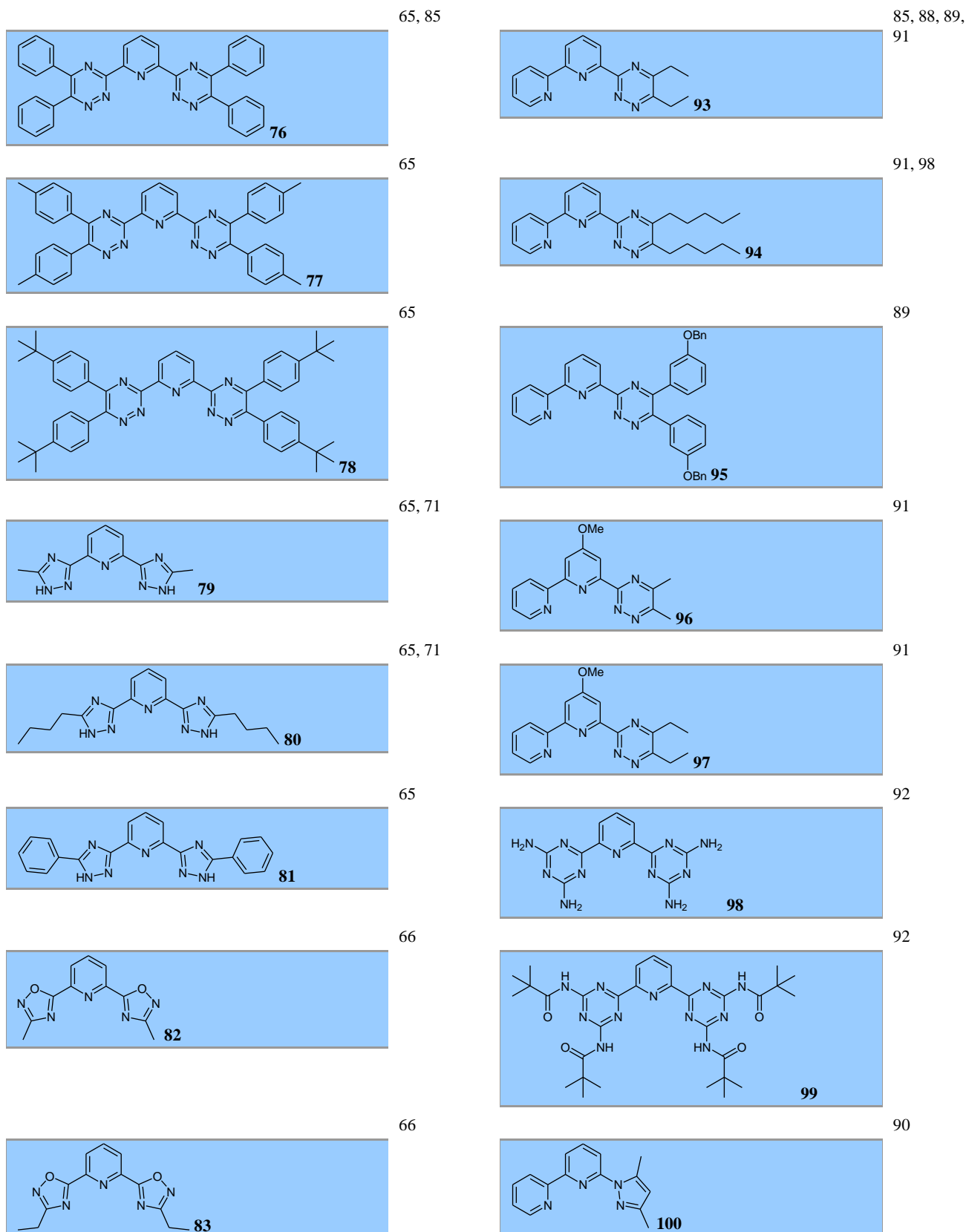
Further solvent extraction studies were carried out on the most promising BTPs; the dipropyl- and diisopropyl-substituted BTPs **70** and **71** (**Table 6**). The dependence of D_{Am} and D_{Eu} on the ligand concentrations showed that the metals were extracted as trisolvates of stoichiometry $[\text{ML}_3](\text{NO}_3)_3$.⁶⁷ The rate of Am(III) extraction was determined by the rate of complexation of the metal, and increased linearly with [BTP] but decreased as $[\text{HNO}_3]$ increased, due to competing protonation of the BTP.⁶⁸ Initial attempts to isolate and structurally characterize the 1:3 complexes of BTP ligands with the lanthanide nitrates were unsuccessful. Only complexes in which the cation contained one or two BTP ligands were obtained,⁶⁹ including an unusual 12-coordinate complex with Nd and Pr.⁷⁰ Structural and theoretical studies with the related bistriazolylpyridine ligands **79** and **80** showed that only complexes containing one ligand and one or more nitrates in the metals inner coordination sphere were formed.⁷¹

Eventually, the X-ray crystal structures of the 1:3 complexes of dipropyl-BTP **70** with Sm(III), Tm(III) and Yb(III) nitrates were obtained and showed that the metal ions were completely enclosed by three ligands while the nitrates were displaced to the outer coordination sphere of the metal.⁷² This provided the first structural evidence for the type of complexes that were involved in the extraction processes. Similarly, BTP complexes of 1:3 stoichiometry were also observed with La(III), Ce(III) and U(III) iodides and triflates.⁷³ The formation of highly lipophilic 1:3 complexes with the BTPs in the presence of nitrate ions, rather than the 1:1 or 1:2 complexes observed with related terdentate *N*-donor ligands, is likely to be responsible for the greatly improved extraction by this remarkable class of ligands. Quantum mechanics calculations suggest that it will be possible to tune the affinity (and basicity) of the BTP ligands through substituent effects,⁷⁴ although this remains to be studied by solvent extraction.

Numerous studies also point to the formation of 1:3 complexes by the BTPs in solution. Electrospray ionization mass spectrometry and time-resolved laser-induced luminescence studies of the complexation of BTPs with Ln(III) in acidic $\text{H}_2\text{O}/\text{MeOH}$ solutions show the presence of stable $\text{Ln}(\text{BTP})_3^{3+}$ complexes. In the more relevant diluent 1-octanol, complexes of 1:1, 1:2 and 1:3 stoichiometry are observed in different relative abundances depending on the lanthanide and the ligand:metal ratio used.⁷⁵ In comparison, Am(III), Cm(III) and Pu(III) exclusively form 1:3 complexes with the BTPs which are thermodynamically more stable than the 1:3 complexes formed with

Table 6 BTP and related ligands studied for their ability to extract and separate An(III) from Ln(III)

Ligand Structure	References	Ligand Structure	References
	70, 78		66
	65, 69, 73, 74, 75a, 75b, 75c, 76b, 77, 78, 85, 91		66
	65, 69, 82b, 84, 85, 91		66
	56, 65, 67, 68, 72, 73, 75, 76a, 76c, 76d, 77, 98		66
	67, 75a, 75b, 75c, 79, 80		66
	65, 82a		66
	65, 69		66
	74, 79, 117		66
	117		85, 88, 91



Eu(III).^{76,56} The difference in stability is mainly enthalpic, pointing to a stronger bonding interaction of the BTPs with An(III) than with Ln(III).

Unusual X-ray crystallographic structures of both 1:1

and 1:2 BTP complexes with the linear uranyl (UO_2^{2+}) cation have been reported.⁷⁷ DFT calculations for the 1:1 complexes showed that the bonding energy in the corresponding TERPY complex was more negative than the BTP complex, and TERPY made a greater orbital contribution to the bonding than BTP.

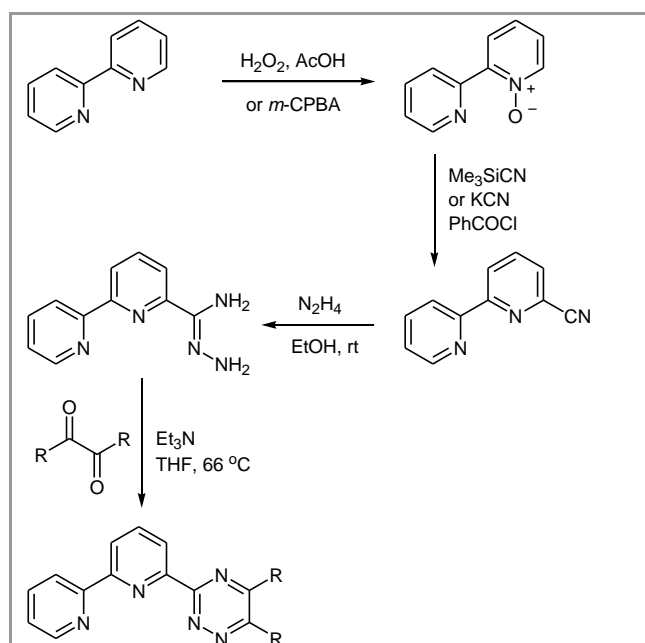
However, the nature of the orbitals involved differed between the two complexes. In the TERPY complex, the bonding was essentially between the UO_2^{2+} ion and the NO_3^- ion, or between the UO_2^{2+} ion and TERPY, whereas in the BTP complex a molecular orbital interaction involving the UO_2^{2+} ion, NO_3^- ions and the triazine rings is found. In this orbital, the lone pair of the nitrogen adjacent to the ligating nitrogen of the triazine ring contributes to the bonding, suggesting that this nitrogen plays an important role in the high selectivity of the BTPs towards An(III). For a series of terdentate nitrogen heterocyclic ligands, *ab initio* calculations of trivalent lanthanide complexes show that complex stability increases as the electron donor ability of the central nitrogen increases and as the electron acceptor ability of the lateral nitrogens increases.⁷⁸ A thermodynamic model has been developed which suggests that, for 4-amino-bis(2,6-(2-pyridyl))-1,3,5-triazine (ADPTZ) **35**, the complexation with Ln(III) is enthalpy driven whereas for BTPs, the complexation is entropy driven.

The behavior of BTPs in their neutral, protonated and complexed forms at oil/water interfaces were investigated using molecular dynamics simulations. It was found that the neutral ligands were weakly surface active, whereas the protonated ligands strongly adsorbed at the interfaces with neutral or acidic water.⁷⁹ Interestingly, the NH^+ group of the protonated ligands point towards the oil phase, rather than the aqueous phase. It appears that complexation occurs right at the interface via the protonated ligand, explaining why the extraction of Am(III) and Eu(III) increases as $[\text{HNO}_3]$ increases. The 1:3 complexes $\text{Eu}(\text{BTP})_3^{3+}$ were also highly surface active, and attracted by water at the interface. The role of synergists such as long chain carboxylic acids and dicarbollide anions appears to be to promote the formation and extraction of the 1:3 complexes adsorbed at the interface by attracting the Eu^{3+} and BTPH^+ cations which would otherwise repel each other.⁸⁰

Laboratory scale SANEX processes were designed and successfully implemented using the dipropyl- and diisopropyl-substituted BTPs **70** and **71** (Table 6).⁸¹ Promising results have also been obtained in the separation of Am(III) from Ln(III) by BTPs immobilized on silica-based resins and supported liquid membranes.⁸² However, in further process development studies it was found that the dialkyl-substituted BTPs were susceptible to acid-catalyzed hydrolysis and radiolytic degradation leading to a decrease in the extraction performance of the diluent.⁸³ This ultimately renders these ligands unsuitable for industrial use, although it has been shown that the addition of nitrobenzene to the organic phase appears to increase the radiolytic stability of the BTPs by acting as a radical scavenger.⁸⁴

At this point, multiple parallel synthesis of a library of 1,2,4-triazine-containing *N*-heterocycles was carried

out in order to identify potentially new ligands for the partitioning of An(III) from Ln(III).⁸⁵ The ligands were synthesized by condensing amide hydrazides with 1,2-diketones. The amide hydrazides were prepared by the reaction of hydrazine with either a 2-cyanopyridine or dithiooxamide.⁸⁶ The 2-cyanopyridines were synthesized by oxidation of the parent pyridines using hydrogen peroxide in acetic acid, followed by a modified Reissert-Henze reaction of the resulting pyridine *N*-oxides.⁸⁷ The best results were observed for the tridentate BTPs and the 6-(1,2,4-triazin-3-yl)-2,2'-bipyridines (hemi-BTPs), in which one of the triazine rings of the BTP has been replaced by a pyridine ring (Scheme 7). None of the bidentate ligands showed any useful extraction properties or Am/Eu selectivity. The hemi-BTP ligands were subsequently chosen for further study.



Scheme 7 Synthesis of 6-(1,2,4-triazin-3-yl)-2,2'-bipyridine ligands

It was found that the hemi-BTPs have properties intermediate between those of the 2,2':6',2''-terpyridines and the BTPs and resemble more closely the 2,2':6',2''-terpyridine ligands.⁸⁸ Although useful selectivities were observed ($\text{SF}_{\text{Am/Eu}} \leq 30$), diethyl hemi-BTP **93** (Table 6) could only extract from nitric acid solutions of low acidity ($[\text{HNO}_3] \leq 0.1 \text{ M}$) and in the presence of the synergist 2-bromodecanoic acid. Similarly, di(benzyloxyphenyl) hemi-BTP **95** selectively extracted Am(III) only from weakly acidic aqueous solutions into octanol or nitrobenzene in synergistic combination with chlorinated cobalt dicarbollide (COSAN) which acts as a hydrophobic anion source.⁸⁹ The formation of both 1:1 and 1:2 complexes of the hemi-BTPs with lanthanide(III) nitrates was observed by ^1H NMR titrations and X-ray crystallography but, unlike the BTPs, there was no evidence for the formation of the more hydrophobic

1:3 complexes. In addition, the ligands were susceptible to decomposition via hydrolysis of the triazine ring in a mixture of nitrous acid and nitric acid.⁸⁸ Replacing the 1,2,4-triazine ring of the hemi-BTPs with a pyrazol-1-yl ring does not improve the extraction properties, and the resulting ligand **100** only extracts Am(III) and Eu(III) from weakly acidic solutions with modest ($SF_{Am/Eu} \leq 8$) selectivity.⁹⁰

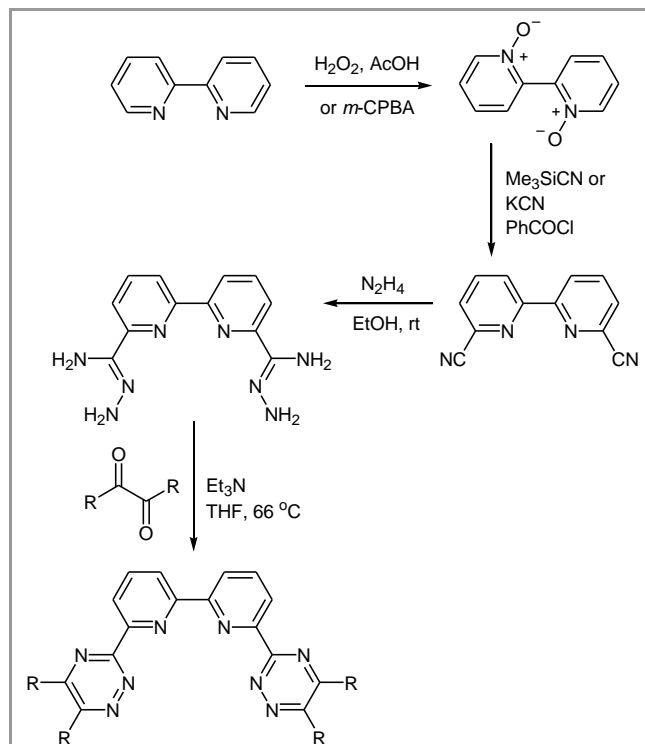
The extraction of various transition metals such as Mn(II), Ni(II), Sr(II), Zr(IV), Ru(II), Cd(II), Pb(II) and Ba(II) by the BTP and the hemi-BTP ligands was investigated since many of these elements are present in the raffinate produced by the PUREX process. For BTP **69** and hemi-BTP **94**, many of these metals were extracted into the organic phase with *D* values in some cases higher than that observed for Am(III).⁹¹ This is significant since these metals in high enough concentrations could disrupt an Am(III) extraction process by sequestering the extractant. A number of homoleptic octahedral complexes of the hemi-BTPs with Mn(II), Fe(II), Co(II) and Zn(II) were also characterized by X-ray crystallography.

Terdentate ligands possessing two outer 1,3,5-triazine rings were then studied to determine if the presence of two adjacent nitrogen atoms in the 1,2,4-triazine rings of the BTPs was responsible for the superior properties of the BTP ligands. The 2,6-bis(1,3,5-triazin-2-yl)pyridines **98** and **99** were synthesized by the reaction of pyridine-2,6-dicarbonitrile and dicyandiamide, followed by acylation of the amino-groups with pivaloyl chloride (in the case of **99**). In solvent extraction experiments, the performance of **98** and **99** was considerably worse than the BTPs under similar conditions either in the presence or absence of a synergist.⁹² Very low distribution ratios and only moderate separation factors ($SF_{Am/Eu} \leq 6$) were observed. This study suggested that the unique properties of the BTPs are indeed related to the presence of two adjacent nitrogen atoms in the triazine rings.

8 BTBP and Related Ligands

The susceptibility of the BTP ligands to hydrolysis and radiolysis prompted us to investigate analogous quadridentate ligands. Process development studies on the BTPs also revealed that the back-extraction (stripping) of An(III) from a loaded organic phase could not be accomplished. This was probably due to the ability of three BTP ligands to enclose completely the inner coordination sphere of the metal. On the other hand, a quadridentate bis-triazine ligand would be expected to have a weaker ligand field and perhaps form less stable 1:1 and/or 1:2 complexes which would leave sufficient room for water, nitrate or stripping agents to coordinate to the metal and displace the ligands. The 6,6'-bis(1,2,4-triazin-3-yl)-2,2'-bipyridine (BTBP) ligands were subsequently investigated.

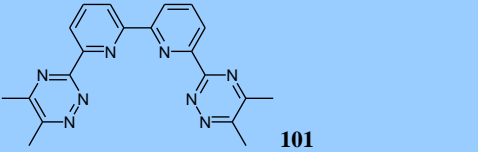

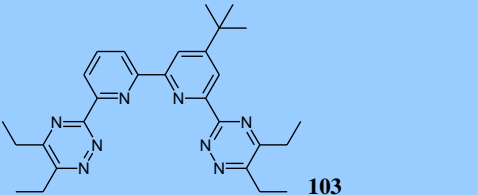
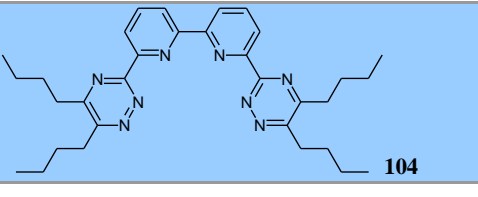
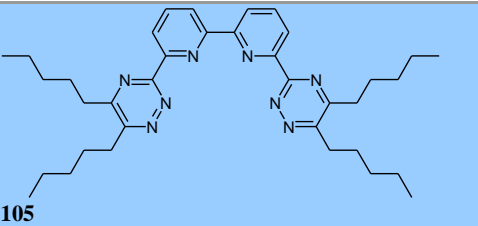
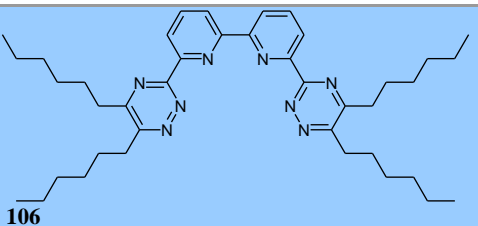
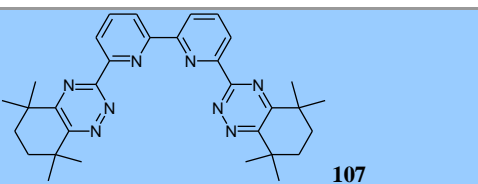
The BTBP ligands are typically synthesized as shown in **Scheme 8**. Oxidation of 2,2'-bipyridine with hydrogen peroxide in acetic acid⁹³ and Reissert-Henze reaction of the resulting bis-*N*-oxide affords the dinitrile.⁹⁴ The addition of hydrazine to the dinitrile generates the diamide dihydrazide⁶⁴ which is condensed with a 1,2-diketone to form the BTBP. BTBPs have previously been employed as synthetic precursors to substituted quaterpyridine ligands.⁶⁴ The BTBP and related ligands that were synthesized and tested for their ability to selectively extract Am(III) from Eu(III) are shown in **Table 7**.

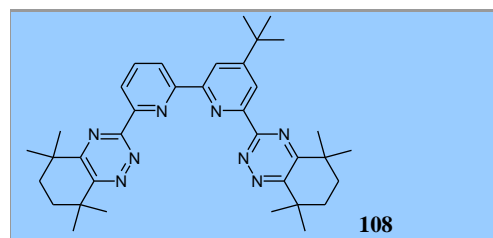


Scheme 8 Synthesis of 6,6'-bis(1,2,4-triazin-3-yl)-2,2'-bipyridine ligands

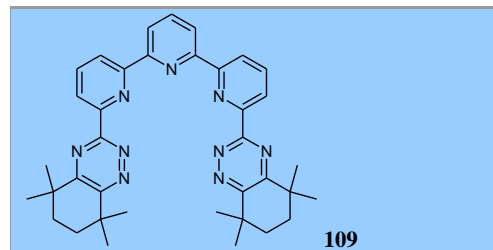
Initial studies with diethyl-BTBP **102** showed promising results. Both high affinity for Am(III) ($D_{Am} = 650 \pm 33$) and high selectivity for Am(III) over Eu(III) ($SF_{Am/Eu} = 160 \pm 16$) were observed in the extraction of the metals from 1 M HNO₃ by a 0.068 M solution of **102** in 1,1,2,2-tetrachloroethane.⁹⁵ The somewhat slow kinetics of extraction could be accelerated using phase-transfer reagents.⁹⁶ Lanthanide nitrate complexes of formula [LaL(NO₃)₃] were synthesized and characterized by X-ray crystallography. The structures showed the lanthanides were coordinated to one planar tetradentate BTBP ligand and three bidentate nitrate ions in a 10-coordinate environment of C₂ symmetry.⁹⁷ No solid state evidence was observed for the formation of 1:2 complexes but NMR titrations with La(NO₃)₃ and Y(NO₃)₃ showed the presence of

Table 7 BTBP and related ligands studied for their ability to extract and separate An(III) from Ln(III)

Ligand Structure	References
	109b, 117
	95, 96, 97, 102, 103, 111, 118
	110, 111, 118
	97, 102
	97, 98, 99, 100, 101, 102, 103, 107, 116, 118, 119
	97, 102
	77, 97, 98, 99, 102, 104, 105, 106, 107, 108, 109, 111, 118, 119, 121



111, 118, 120



112

both 1:1 and 1:2 complexes in solution. Further evidence for the formation of complexes of these stoichiometries was obtained from spectrophotometric titrations of dipentyl-BTBP **105** and CyMe₄-BTBP **107** with La(III), Eu(III) and Yb(III) in methanol.⁹⁸ However, only the 1:2 complexes of CyMe₄-BTBP **107** and dipentyl-BTBP **105** with Am(III) in various diluents were observed by electrospray ionization mass spectrometry.⁹⁹

Further detailed solvent extraction studies were carried out on more lipophilic BTBPs bearing longer alkyl side chains. Dipentyl-BTBP **105** showed variable distribution ratios but high separation factors in the extraction of Am(III) and Eu(III) into a range of organic diluents which suggested that back extraction of the metals would be possible through careful choice of diluent and by adjusting the acidity of the aqueous phase.¹⁰⁰ The most suitable diluent was cyclohexanone and rapid extraction kinetics were observed in this diluent (equilibrium was reached within 5 minutes).¹⁰¹ Another feature of the BTBPs is that, like the BTPs, the extraction of both Am(III) and Eu(III) increases as [HNO₃] increases.

The thermodynamics of the extraction showed that Am(III) had a more negative Gibbs free energy of extraction than Eu(III). Back extraction of Am(III) using glycolic acid and recycling of the organic phase were successfully demonstrated showing that, unlike the BTP ligands, a continuous extraction process was feasible. However, **105** was sensitive to radiolysis and *D*_{Am} decreased by ca. 80 % on exposure of the diluent to ca. 17 kGy of gamma radiation.¹⁰⁰

The extraction of Am(III), Th(IV), Np(V) and U(VI) by a range of BTBP ligands has been investigated to determine the effect of the oxidation state of the actinide ions on their extraction.¹⁰² It was found that the extraction increased in the order Am(III) < Th(IV) < Np(V) < U(VI) at low BTBP concentrations but increased in the order Th(IV) < U(VI) < Np(V) <

Am(III) at high BTBP concentrations. This study suggests that the trivalent and pentavalent actinides can be separated from the tetravalent and hexavalent actinides using the BTBPs. In addition, the BTBPs also extracted several D-block metals to a significant extent, in common with the terdentate BTP and hemi-BTP ligands.^{100,103}

The most promising of the BTBP ligands is CyMe₄-BTBP **107**, which is designed to be more resistant to radiolysis (this will be discussed in detail later). This ligand is the most studied of the BTBP ligands, and its solvent extraction chemistry has been studied in detail.¹⁰⁴ High selectivities were observed for Am(III) and Cm(III) over the entire lanthanide series including yttrium. When dissolved in 1-octanol (0.02 M), the ligand gave $D_{\text{Am}} = \text{ca. } 4.5$ and $SF_{\text{Am/Eu}} = \text{ca. } 140$ in the extraction of Am(III) and Eu(III) from 0.5 M HNO₃, and equilibrium was reached within 60 minutes. The slow kinetics of extraction was considerably improved by the addition of DMDOHEMA **4** (Figure 1) to the organic phase, and equilibrium was reached after only 5 minutes. The dependence of D_{Am} on [CyMe₄-BTBP] showed the metals were extracted as 1:2 complexes. The ligand showed excellent hydrolytic stability on long term (2 months) exposure to 1 M HNO₃ with no reduction in D values or $SF_{\text{Am/Eu}}$. In addition, the selective back-extraction of the lanthanides (scrubbing), followed by the actinides (stripping) was accomplished using aqueous solutions of 0.5 M HNO₃ and 0.5 M glycolic acid at pH 4, respectively.

After some optimization of the conditions,¹⁰⁵ a laboratory scale SANEX process was successfully demonstrated on a genuine spent fuel solution using CyMe₄-BTBP **107** and DMDOHEMA **4** dissolved in 1-octanol.¹⁰⁶ The counter current extraction process was carried out in centrifugal contactors and comprised 16 stages (9 stages for extraction, 3 stages for lanthanide back-extraction and 4 stages for actinide back-extraction). More than 99.9 % of the Am(III) and Cm(III) were recovered from the 2 M HNO₃ feed solution and more than 99.9 % of the lanthanides were directed to the raffinate.

CyMe₄-BTBP **107** has also been investigated as a potential ligand for a GANEX (Group ActiNide EXtraction) separation process. In this process, it is planned to selectively extract the actinides directly from the PUREX raffinate, removing the need for a preceding DIAMEX process. Under these conditions, the actinides will need to be extracted in different oxidation states and the metal loading of the aqueous phase will be far higher than in a SANEX process. Under process-like conditions, solutions of CyMe₄-BTBP **107** and tributyl phosphate **1** (used in the PUREX process) in cyclohexanone selectively extracted Am, Pu, Np and U from 4 M nitric acid solutions of high metal loading (> 9000 ppm),¹⁰⁷ suggesting that a GANEX process is feasible. It was shown that each ligand extracted independently of the other and no synergistic effects were observed.

Recently, the first crystal structures of actinide complexes of CyMe₄-BTBP **107** have been reported. Treatment of the ligand with the uranyl salts UO₂(OTf)₂ and UO₂I₂(THF)₃ under anhydrous conditions led to the formation of mononuclear 1:1 complexes in which the ligand coordinates to the metal in a tetradentate fashion perpendicular to the plane of the UO₂²⁺ ion.¹⁰⁸ In the presence of moist air, the complexes crystallized as the dinuclear complexes [(UO₂L)–O–(UO₂L)²⁺] containing a flexible μ -oxo bridge. Mononuclear 1:1 complexes of **107** were also obtained with tetravalent Cp^{*}₂U²⁺ ions, and with pentavalent UO₂⁺ ions which were formed by reduction of the corresponding UO₂²⁺ complexes with KC₅Me₅ or TiC₅H₅.¹⁰⁹ A very unusual dinuclear μ -oxo mixed valence U(IV)/U(VI) complex was isolated from a mixture of U(OTf)₄, UO₂(NO₃)₂ and CyMe₄-BTBP **107**.⁷⁷ In this complex, two ligands **107** coordinate to the U(IV) ion in an orthogonal arrangement while the nitrates coordinate to the U(VI) ion.

A comparative investigation of the complexation of Cm(III) and Eu(III) by ¹Bu-C2-BTBP **103** using time-resolved laser-induced fluorescence spectroscopy (TRLFS) was carried out under conditions relevant to the extraction process. Complexes of 1:2 stoichiometry were observed for both metal ions but the stability constant for the Cm(III) complex was found to be approx. two orders of magnitude higher than that for the Eu(III) complex.¹¹⁰ The enthalpy of complexation of Cm(III) was more negative than that of Eu(III), while the entropy of complexation of both metal ions was very similar. This study showed that the remarkably high selectivity of the BTBPs for An(III) over Ln(III) was driven by the greater thermodynamic stability of its complexes with An(III).

Recent quantum mechanics calculations show that the BTBPs are less basic, and afford complexes that are less thermodynamically stable than those of their rigidified phenanthroline (BTP_{phen}) analogues. Molecular dynamics simulations at the oil/water interface reveal that CyMe₄-BTBP **107** is only surface active when protonated, and when complexed to a single Eu(III) cation.¹¹¹ As with the BTPs, the NH⁺ group of the protonated BTBP ligand **107** orients away from the aqueous phase to avoid hydrophobic contacts with the aliphatic rings. These results suggest that the extraction of Am(III) and Eu(III) by **107** takes place at the interface via the protonated ligand in acidic conditions, and explains why the extraction kinetics is slow in the absence of a phase transfer agent.

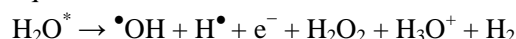
In an effort to better understand the origins of the high selectivity of bis-triazine ligands for An over Ln, we investigated the pentadentate ligand CyMe₄-BTTP **109**, which contains an additional pyridine ring compared to the BTBPs. Although mildly selective ($SF_{\text{Am/Eu}} \leq 12$), the ligand was essentially unable to

extract Am(III) or Eu(III) from HNO₃ into 1-octanol in the absence of a synergist ($D_{\text{Am}} < 0.04$), making it less efficient than the tetradentate CyMe₄-BTBP **107** by about two orders of magnitude.¹¹² In nitrate media, only the 1:1 complexes with Ln(III) were observed by NMR and mass spectrometry, whereas the more hydrophobic 1:2 complexes were formed in perchlorate media, indicating that nitrate ions competed with the ligand **109** for coordination sites on the metal resulting in less hydrophobic 1:1 complexes which were not easily extracted.

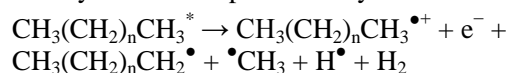
9 Improving Radiolytic Stability through Ligand Design

In a potential SANEX process the organic and aqueous phases will be exposed to high levels of ionizing radiation due to the presence of radioactive α - and γ -emitting radionuclides. This will impact the performance of any solvent extraction processes used for nuclear fuel reprocessing, and a knowledge of these effects is critical for the design of robust solvent systems. The effect of radiation chemistry on solvent extraction processes was recently reviewed.¹¹³ It has been estimated that the organic phase in a SANEX process will receive a total annual dose of between 100 kGy and 1 MGy, depending on the type of waste being reprocessed.⁸⁴ Direct absorption of this energy by the diluent molecules of the organic and aqueous phases leads to the formation of electronically excited states, free-radical species, solvated electrons and ions which can subsequently react with the extractant molecules.

The radiolysis of water can be represented by the equation:¹¹⁴



while for a typical alkane-like organic diluent, the radiolysis can be represented by:¹¹⁵



Process development studies on the dialkyl-BTPs such as diisopropyl-BTP **71** showed that the ligands were susceptible to indirect radiolysis caused by chemical attack at their reactive benzylic positions by free-radical species, with detrimental effects on their extraction ability and selectivity.^{83b} Furthermore, the extraction performance of dipentyl-BTBP **105** decreased with increasing doses of γ -radiation, and the likely degradation products were identified by mass spectrometry.¹¹⁶ Many of the degradation products arose from hydrogen atom abstraction and subsequent oxidation of the benzylic positions eventually leading to cleavage of the alkyl groups. The replacement of the labile benzylic hydrogens with alkyl groups thus became the primary goal in the design and synthesis of selective actinide extractants with improved resistance to radiolysis.

Subsequently, attempts were made to synthesize a BTP ligand containing *tert*-butyl groups attached to the triazine rings, rather than primary alkyl groups. Unfortunately, the condensation reaction of pivalil with pyridine-2,6-dicarboxamide hydrazone (**Scheme 6** where R = *t*-Bu) failed to afford the desired ligand, probably due to the sterically hindered nature of the α -diketone or its inability to adopt the *cis*-conformation required for the condensation reaction. It was reasoned that a cyclic α -diketone would be more reactive as the carbonyl groups would be locked in the required *cis*-conformation. Gratifyingly, the reaction of pyridine-2,6-dicarboxamide hydrazone with the cyclic α -diketones 3,3,6,6-tetramethylcyclohexane-1,2-dione and 1,1',4,4'-tetramethylnaphthalene-2,3-dione furnished in high yields the annulated ligands CyMe₄-BTP **74** and BzCyMe₄-BTP **75**, respectively (**Table 6**).¹¹⁷

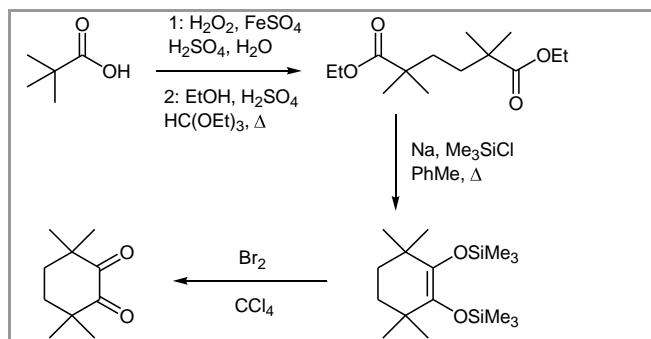
In solvent extraction experiments, CyMe₄-BTP **74** proved to be a more effective extractant than diisopropyl-BTP **71** in the presence of the phase-modifier DMDOHEMA. Typically, for 0.01 M solutions of the ligands in 1-octanol after 15 minutes of phase contact with 0.5 M HNO₃, $D_{\text{Am}} = \text{ca. } 500$ and $\text{SF}_{\text{Am/Eu}} = \text{ca. } 5000$ for CyMe₄-BTP **74** whereas $D_{\text{Am}} = \text{ca. } 30$ and $\text{SF}_{\text{Am/Eu}} = \text{ca. } 400$ for diisopropyl-BTP **71**. For BzCyMe₄-BTP **75**, $D_{\text{Am}} = \text{ca. } 10$ and $\text{SF}_{\text{Am/Eu}} = \text{ca. } 500$ under identical conditions.¹¹⁷ The remarkably high selectivity for Am(III) over Eu(III) of CyMe₄-BTP **74** was far in excess of what was observed with other *N*-donor ligands. The formation of the expected 1:3 complexes of **74** with the lanthanides was confirmed by ¹H NMR titrations and X-ray crystallography.

In contrast to other BTPs, both of the annulated BTPs **74** and **75** were stable to hydrolysis and were not hydrolyzed by boiling nitric acid (3 M) over 24 hours. After exposure of 1-octanol solutions of the ligands to 100 kGy of γ -radiation, there was significant degradation of CyMe₄-BTP **74** and diisopropyl-BTP **71** (ca. 80 %), but no degradation of BzCyMe₄-BTP **75** occurred. It is thought that the additional aromatic rings present in **75** compared to **74** are able to dissipate some of the energy absorbed by the ligand during direct radiolysis. Theoretical simulations of electron loss and gain on various fragments of **74** and **75** suggest that, upon electron loss (as would occur during direct radiolysis), most of the charge resides in the benzene rings of **75** (64 %) but in the triazine rings of **74** (45 %). This view is supported by frontier orbital calculations which show that there are two pairs of doubly-degenerate HOMOs localized on the outer benzene rings of **75** which are not present in **74**.¹¹⁷

The stability of the tetradentate ligands CyMe₄-BTBP **107** and its more soluble¹¹⁸ derivative MF2-BTBP **108** towards radiolysis was then investigated. When exposed to low doses of gamma irradiation (< 30 kGy), CyMe₄-BTBP **107** dissolved in hexanol or

cyclohexanone showed good stability and the extraction of Am(III) and Eu(III) remained largely unaffected. In contrast, the extraction efficiency of dipentyl-BTBP **105**, which has benzylic hydrogens, decreased as the total absorbed dose increased under the same conditions.¹¹⁹ However, the resistance of MF2-BTBP **108** dissolved in cyclohexanone towards gamma radiolysis was found to depend on the dose rate. At low dose rates (ca. 15 Gy/h), the extraction of Am(III) and Eu(III) by MF2-BTBP **108** was more or less unchanged up to an absorbed dose of 100 kGy but at high dose rates (ca. 1.2 kGy/h), the extraction of Am(III) and Eu(III) decreased with increasing dose, indicating that secondary processes (indirect radiolysis) were likely to be responsible for the degradation rather than the radiation directly.¹²⁰ In a similar study, CyMe₄-BTBP **107** in 1-octanol was more sensitive to gamma radiolysis (0.22 kGy/h) than to alpha radiolysis (1 kGy/h), and the rate of degradation of **107** was 40 % higher when exposed to gamma radiation.¹²¹

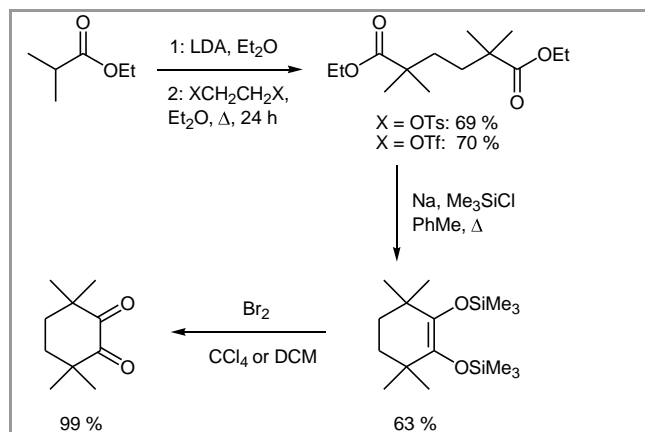
One of the drawbacks of CyMe₄-BTP **74**, CyMe₄-BTBP **107** and MF2-BTBP **108** is the difficulties encountered in synthesizing large quantities of the ligands for process development studies. This is due to the relatively inefficient method of synthesizing the α -diketone precursor (Scheme 9). The first step involves the generation of 2,2,5,5-tetramethyladipic acid by the free-radical dimerization of pivalic acid mediated by Fenton's reagent (H₂O₂ and FeSO₄).¹²² This reaction is reported to give yields of ca. 37 % but in our hands, yields of 15 % or lower were obtained. The diacid is then esterified and the diester is cyclized by an intramolecular acyloin reaction using sodium and chlorotrimethylsilane in toluene. Oxidation of the enediolate bis-silyl ether with bromine then affords 3,3,6,6-tetramethylcyclohexane-1,2-dione. The acyloin reaction may also be carried out in the absence of chlorotrimethylsilane generating the α -hydroxyketone which is then oxidised to the α -diketone using copper(II) acetate.¹²³



Scheme 9 Synthesis of 3,3,6,6-tetramethylcyclohexane-1,2-dione

Although suitable for synthesizing small quantities of **74**, **107** and **108**, this synthetic route is impractical for large-scale synthesis and thus presents a barrier to the application of the ligands in an industrial SANEX

process. Recently, a new and more efficient synthesis of 3,3,6,6-tetramethylcyclohexane-1,2-dione was developed (Scheme 10) which by-passes the problematic first step in the previous method.¹²⁴ Alkylation of the enolate of ethyl isobutyrate with 1,2-disulfonate esters derived from ethylene glycol afforded the diester intermediate in good yields (69–70 %).



Scheme 10 Modified and improved synthesis of 3,3,6,6-tetramethylcyclohexane-1,2-dione

Interestingly, when this reaction was conducted using softer 1,2-dihalides as electrophiles, mixtures of products were obtained of which the major product was the corresponding ethyl α -haloisobutyrate. This product was formed by attack of the enolate on the halogen atom with subsequent elimination of ethylene.¹²⁵ Completion of the synthesis as before (intramolecular acyloin reaction/oxidation) yielded the α -diketone with significantly higher yields than the previous method (44 % overall compared to < 5 % overall). This facile synthesis of the diketone precursor to CyMe₄-based ligands such as **74**, **107** and **108** removed one of the biggest obstacles to the development of an industrial SANEX process, namely, the ability to synthesize the ligands on a large scale.

Conclusions and Outlook

In summary, we have surveyed recent advances in Europe in the design, synthesis, study and development of soft N- and S-donor ligands for separating actinides from lanthanides in the nuclear fuel cycle. Bis-1,2,4-triazine ligands composed only of the elements C, H, O and N have emerged as the most promising family of ligands for this purpose, showing levels of selectivity between the chemically similar actinides and lanthanides once thought impossible to achieve. These reagents and their extracted metal complexes possess a high enough range of solubilities in suitable diluents such as 1-octanol for use in the SANEX process. Optimum levels of affinity that permit back-extraction of the

metals have been achieved with the BTBP ligands. The BTBPs are not susceptible to hydrolysis and are resistant to radiolysis at low dose rates. The ease of synthesis of these reagents has been recently improved, and CyMe₄-BTBP **107** is now commercially available. Furthermore, the ligands are capable of extracting from solutions of high acidity.

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References

- (1) (a) *World Energy Outlook*, International Atomic Energy Agency (IAE/OECD), **2007**. (b) *Nuclear Energy Outlook 2008*, OECD/NEA, Paris, **2008**.
- (2) *Implications of partitioning and transmutation in radioactive waste management*, International Atomic Energy Agency, Technical Report Series, no. 435, Vienna, **2004**.
- (3) (a) McKibben, J. M. *Radiochim. Acta* **1984**, *36*, 3. (b) Sood, D. D.; Patil, S. K. *J. Rad. Nucl. Chem.* **1996**, *203*, 547. (c) Musikas, C.; Schultz, W.; Liljenzin, J.-O. in *Solvent Extraction Principles and Practice*, 2nd ed; Rydberg, J.; Cox, M.; Musikas, C.; Choppin, G., Eds.; Marcel Dekker, Inc.: New York, **2004**, pp 507–557. (d) Nash, K. L.; Madic, C.; Mathur, J.; Lacquement, J. in *The Chemistry of the Actinide and Transactinide Elements*, 3rd ed; Katz, J. J.; Morss, L. R.; Edelstein, N. M.; Fuger, J., Eds.; Springer: Dordrecht, **2006**, pp 2622–2798.
- (4) Magill, J.; Berthou, V.; Haas, D.; Galy, J.; Schenkel, R.; Wiese, H.-W.; Heusener, G.; Tommasi, J.; Youinou, G. *Nuclear Energy* **2003**, *42*, 263.
- (5) (a) *Actinide and fission product partitioning and transmutation status and assessment report*, OECD/NEA, Paris, **1999**. (b) Salvatores, M.; *Nucl. Eng. Design* **2005**, *235*, 805. (c) Salvatores, M.; Palmiotti, G. *Progress in Particle and Nuclear Physics* **2011**, *66*, 144.
- (6) (a) Nash, K. L. *Solvent Extr. Ion Exch.* **1993**, *11*, 729. (b) Mathur, J. N.; Murali, M. S.; Nash, K. L. *Solvent Extr. Ion Exch.* **2001**, *19*, 357.
- (7) (a) Cuillerdier, C.; Musikas, C. *Sep. Sci. Technol.* **1991**, *26*, 1229. (b) Modolo, G.; Vijgen, H.; Serrano-Purroy, D.; Christiansen, B.; Malmbeck, R.; Sorel, C.; Baron, P. *Sep. Sci. Technol.* **2007**, *42*, 439.
- (8) Horwitz, E. P.; Kalina, D. G.; Diamond, H.; Vandegrift, G. F.; Schulz, W. W. *Solvent Extr. Ion Exch.* **1985**, *3*, 75.
- (9) (a) Cuillerdier, C.; Musikas, C.; Hoel, P.; Nigond, L.; Vitart, X. *Sep. Sci. Technol.* **1991**, *26*, 1229. (b) Nigond, L.; Condamines, N.; Cordier, P. Y.; Livet, J.; Madic, C.; Cuillerdier, C.; Musikas, C.; Hudson, M. J. *Sep. Sci. Technol.* **1995**, *30*, 2075.
- (10) Chan, G. Y. S.; Drew, M. G. B.; Hudson, M. J.; Iveson, P. B.; Liljenzin, J.-O.; Skälberg, M.; Spjuth, L.; Madic, C. *J. Chem. Soc. Dalton Trans.* **1997**, 649.
- (11) Serrano-Purroy, D.; Baron, P.; Christiansen, B.; Malmbeck, R.; Sorel, C.; Glatz, J.-P. *Radiochim. Acta* **2005**, *93*, 351.
- (12) (a) Modolo, G.; Asp, H.; Schreinemachers, C.; Vijgen, H. *Solvent Extr. Ion Exch.* **2007**, *25*, 703. (b) Magnusson, D.; Christiansen, B.; Glatz, J.-P.; Malmbeck, R.; Modolo, G.; Serrano-Purroy, D.; Sorel, C. *Solvent Extr. Ion Exch.* **2009**, *27*, 26.
- (13) (a) Cotton, S. in *Comprehensive Coordination Chemistry II*, Vol. 3; McCleverty, J. A.; Meyer, T. J., Eds.; Elsevier: Oxford, **2004**, pp 93–188. (b) Burns, C. J.; Neu, M. P.; Boukhalfa, H.; Gutowski, K. E.; Bridges, N. J.; Rogers, R. D. in *Comprehensive Coordination Chemistry II*, Vol. 3; McCleverty, J. A.; Meyer, T. J., Eds.; Elsevier: Oxford, **2004**, pp 189–332. (c) Katz, J. J.; Morss, L. R.; Edelstein, N. M.; Fuger, J. in *The Chemistry of the Actinide and Transactinide Elements*, Vol. 1; Katz, J. J.; Morss, L. R.; Edelstein, N. M.; Fuger, J., Eds.; Springer: Dordrecht, **2006**, pp 1–17.
- (14) Nilsson, M.; Nash, K. L. *Solvent Extr. Ion Exch.* **2007**, *25*, 665.
- (15) Hudson, M. J.; *Czech. J. Phys.* **2003**, *53*, A305.
- (16) (a) Dam, H. H.; Reinhoudt, D. N.; Verboom, W. *Chem. Soc. Rev.* **2007**, *36*, 367. (b) Ekberg, C.; Fermvik, A.; Retegan, T.; Skarnemark, G.; Foreman, M. R. S.; Hudson, M. J.; Englund, S.; Nilsson, M. *Radiochim. Acta* **2008**, *96*, 225. (c) Kolarik, Z. *Chem. Rev.* **2008**, *108*, 4208. (d) Hill, C. in *Ion Exchange and Solvent Extraction: A Series of Advances*, Vol. 19; Moyer, B. A., Ed.; CRC Press: Boca Raton, **2010**, pp 119–194.
- (17) (a) Zhu, Y. *Radiochim. Acta* **1995**, *68*, 95. (b) Zhu, Y.; Chen, J.; Jiao, R. *Solvent Extr. Ion Exch.* **1996**, *14*, 61. (c) Chen, J.; Zhu, Y.; Jiao, R. *Sep. Sci. Technol.* **1996**, *31*, 2723. (d) Hill, C.; Madic, C.; Baron, P.; Ozawa, M.; Tanaka, Y. *J. Alloys Compd.* **1998**, *271–273*, 159. (e) Wang, X.; Zhu, Y.; Jiao, R. *J. Rad. Nucl. Chem.* **2002**, *251*, 487.
- (18) Ionova, G.; Ionov, S.; Rabbe, C.; Hill, C.; Madic, C.; Guillaumont, R.; Krupa, J. C. *Solvent Extr. Ion Exch.* **2001**, *19*, 391.
- (19) (a) Bhattacharyya, A.; Mohapatra, P. K.; Manchanda, V. K. *Solvent Extr. Ion Exch.* **2006**, *24*, 1. (b) Bhattacharyya, A.; Mohapatra, P. K.; Manchanda, V. K. *Radiochim. Acta* **2010**, *98*, 141.
- (20) (a) Peterman, D. R.; Law, J. D.; Todd, T. A.; Tillotson, R. D. in *Separations for the Nuclear Fuel Cycle in the 21st Century*, Vol 933; Lumetta, G. J.; Nash, K. L.; Clark, S. B.; Fries, J. I., Eds.; American Chemical Society: Washington DC, **2006**, pp 251–259. (b) Law, J. D.; Peterman, D. R.; Todd, T. A.; Tillotson, R. D. *Radiochim. Acta* **2006**, *94*, 261.
- (21) (a) Wei, Y.; Kumagai, M.; Takashima, Y.; Modolo, G.; Odoj, R. *Nuclear Technology* **2000**, *132*, 413. (b) Bhattacharyya, A.; Mohapatra, P. K.; Manchanda, V. K. *J. Chromatography A* **2006**, *1123*, 26. (c) Bhattacharyya, A.; Mohapatra, P. K.; Manchanda, V. K. *Solvent Extr. Ion Exch.* **2007**, *25*, 27.
- (22) Mimura, H.; Hoshi, H.; Akiba, K.; Onodera, Y. *J. Rad. Nucl. Chem.* **2001**, *247*, 375.
- (23) Hoshi, H.; Tsuyoshi, A.; Akiba, K. *J. Rad. Nucl. Chem.* **2000**, *243*, 621.
- (24) Wang, X.; Zhu, Y.; Jiao, R. *Solvent Extr. Ion Exch.* **2001**, *19*, 1007.
- (25) Modolo, G.; Odoj, R. *J. Rad. Nucl. Chem.* **1998**, *228*, 83.
- (26) (a) Modolo, G.; Odoj, R. *J. Alloys Compd.* **1998**, *271–273*, 248. (b) Modolo, G.; Odoj, R. US5966584 (**1999**). (c) Geist, A.; Weigl, M.; Gompfer, K. *Sep. Sci. Technol.*

- 2002**, 37, 3369.
- (27) Modolo, G.; Odoj, R. *Solvent Extr. Ion Exch.* **1999**, 17, 33.
- (28) Modolo, G.; Seekamp, S. *Solvent Extr. Ion Exch.* **2002**, 20, 195.
- (29) (a) Modolo, G.; Nabet, S. *Solvent Extr. Ion Exch.* **2005**, 23, 359. (b) Modolo, G.; Odoj, R. US20070009410A1 (**2007**).
- (30) Modolo, G.; Kluxen, P.; Geist, A. *Radiochim. Acta* **2010**, 98, 193.
- (31) Ionova, G.; Ionov, S.; Rabbe, C.; Hill, C.; Madic, C.; Guillaumont, R.; Modolo, G.; Krupa, J. C. *New J. Chem.* **2001**, 25, 491.
- (32) (a) Klaehn, J. R.; Peterman, D. R.; Harrup, M. K.; Tillotson, R. D.; Luther, T. A.; Law, J. D.; Daniels, L. M. *Inorg. Chim. Acta* **2008**, 361, 2522. (b) Peterman, D. R.; Greenhalgh, M. R.; Tillotson, R. D.; Klaehn, J. R.; Harrup, M. K.; Luther, T. A.; Law, J. D. *Sep. Sci. Technol.* **2010**, 45, 1711.
- (33) Qichu, X.; Jianfeng, W.; Yinzhong, C.; Lixing, Z.; Yusheng, Y. *Radiochim. Acta* **2008**, 96, 771.
- (34) Weigl, M.; Denecke, M. A.; Panak, P. J.; Geist, A.; Gompfer, K. *Dalton Trans.* **2005**, 1281.
- (35) Coupe, B.; Boehme, C.; Wipff, G. *J. Phys. Chem. B* **2003**, 107, 9484.
- (36) Cordier, P. Y.; Hill, C.; Baron, P.; Madic, C.; Hudson, M. J.; Liljezin, J.-O. *J. Alloys Compd.* **1998**, 271–273, 738.
- (37) Hagström, I.; Spjuth, L.; Enarsson, Å.; Liljezin, J.-O.; Skälberg, M.; Hudson, M. J.; Iveson, P. B.; Madic, C.; Cordier, P. Y.; Hill, C.; Francois, N. *Solvent Extr. Ion Exch.* **1999**, 17, 221.
- (38) Drew, M. G. B.; Iveson, P. B.; Hudson, M. J.; Liljezin, J.-O.; Spjuth, L.; Cordier, P. Y.; Enarsson, Å.; Hill, C.; Madic, C. *J. Chem. Soc. Dalton Trans.* **2000**, 821.
- (39) Grigoriev, M. S.; Den Auwer, C.; Madic, C. *Acta Cryst.* **2001**, C57, 1141.
- (40) Berthet, J.-C.; Rivière, C.; Miquel, Y.; Nierlich, M.; Madic, C.; Ephritikhine, M. *Eur. J. Inorg. Chem.* **2002**, 1439.
- (41) Drew, M. G. B.; Hudson, M. J.; Iveson, P. B.; Russell, M. L.; Liljezin, J.-O.; Skälberg, M.; Spjuth, L.; Madic, C. *J. Chem. Soc. Dalton Trans.* **1998**, 2973.
- (42) (a) Case, F. H.; Koft, E. *J. Am. Chem. Soc.* **1959**, 81, 905. (b) Forsberg, J. H.; Spaziano, V. T.; Klump, S. P.; Sanders, K. M. *J. Het. Chem.* **1988**, 25, 767. (c) Hudson, M. J.; Chan, G. Y. S.; Madic, C.; Baron, P. GB2296917A (**1996**). (d) Drew, M. G. B.; Hudson, M. J.; Iveson, P. B.; Russell, M. L.; Madic, C. *Acta Cryst.* **1998**, C54, 985.
- (43) (a) Chan, G. Y. S.; Drew, M. G. B.; Hudson, M. J.; Isaacs, N. S.; Byers, P. *Polyhedron* **1996**, 15, 3385. (b) Drew, M. G. B.; Hudson, M. J.; Iveson, P. B.; Madic, C. *Acta Cryst.* **2000**, C56, 434.
- (44) (a) Musikas, C. *Actinide-Lanthanide Group Separation Using Sulfur and Nitrogen Donor Extractants*. Proceedings of the International Symposium on Actinide/Lanthanide Separations, Honolulu, Hawaii, USA (**1984**). (b) Madic, C.; Hudson, M. J. *High-Level Liquid Waste Partitioning by Means of Completely Incinerable Extractants*. Final Report, European Commission contract No. F12W-CT91-0112, EUR18038 (**1998**).
- (45) Drew, M. G. B.; Hudson, M. J.; Iveson, P. B.; Madic, C.; Russell, M. L. *J. Chem. Soc. Dalton Trans.* **2000**, 2711.
- (46) Boubals, N.; Drew, M. G. B.; Hill, C.; Hudson, M. J.; Iveson, P. B.; Madic, C.; Russell, M. L.; Youngs, T. G. A. *J. Chem. Soc. Dalton Trans.* **2002**, 55.
- (47) Miguiditchian, M.; Guillauneux, D.; Guillaumont, D.; Moisy, P.; Madic, C.; Jensen, M. P.; Nash, K. L. *Inorg. Chem.* **2005**, 44, 1404.
- (48) Addison, A. W.; Burke, P. J. *J. Het. Chem.* **1981**, 18, 803.
- (49) Drew, M. G. B.; Hill, C.; Hudson, M. J.; Iveson, P. B.; Madic, C.; Vaillant, L.; Youngs, T. G. A. *New J. Chem.* **2004**, 28, 462.
- (50) (a) Andersson, S.; Ekberg, C.; Foreman, M. R. S.; Hudson, M. J.; Liljezin, J.-O.; Nilsson, M.; Skarnemark, G.; Spahiu, K. *Solvent Extr. Ion Exch.* **2003**, 21, 621. (b) Andersson, S.; Ekberg, C.; Liljezin, J.-O.; Nilsson, M.; Skarnemark, G. *Radiochim. Acta* **2004**, 92, 863.
- (51) Wietzke, R.; Mazzanti, M.; Latour, J.-M.; Pécaut, J.; Cordier, P.-Y.; Madic, C. *Inorg. Chem.* **1998**, 37, 6690.
- (52) Ishimori, K.; Watanabe, M.; Kimura, T.; Yaita, T.; Yamada, T.; Kataoka, Y.; Shinoda, S.; Tsukube, H. *Chem. Lett.* **2005**, 34, 1112.
- (53) Wietzke, R.; Mazzanti, M.; Latour, J.-M.; Pécaut, J. *J. Chem. Soc. Dalton Trans.* **2000**, 4167.
- (54) Natrajan, L.; Pécaut, J.; Mazzanti, M.; LeBrun, C. *Inorg. Chem.* **2005**, 44, 4756.
- (55) Bravard, F.; Rosset, C.; Delangle, P. *Dalton Trans.* **2004**, 2012.
- (56) Denecke, M. A.; Panak, P. J.; Burdet, F.; Weigl, M.; Geist, A.; Klenze, R.; Mazzanti, M.; Gompfer, K. C. *Chimie* **2007**, 10, 872.
- (57) (a) Watanabe, M.; Mirvaliev, R.; Tachimori, S.; Takeshita, K.; Nakano, Y.; Morikawa, K.; Mori, R. *Chem. Lett.* **2002**, 31, 1230. (b) Mirvaliev, R.; Watanabe, M.; Matsumura, T.; Tachimori, S.; Takeshita, K. *J. Nucl. Sci. Technol.* **2004**, 41, 1122.
- (58) Matsumura, T.; Takeshita, K. in *Separations for the Nuclear Fuel Cycle in the 21st Century*, Vol 933; Lumetta, G. J.; Nash, K. L.; Clark, S. B.; Friese, J. I., Eds.; American Chemical Society: Washington DC, **2006**, pp 261–273.
- (59) Karmazin, L.; Mazzanti, M.; Gateau, C.; Hill, C.; Pécaut, J. *Chem. Commun.* **2002**, 2892.
- (60) (a) Wibaut, J. P.; de Jonge, A. P.; van der Voort, H. G. P.; Otto, P. Ph. H. L. *Rec. Trav. Chim. Pays-Bas* **1951**, 70, 1054. (b) Hannon, M. J.; Mayers, P. C.; Taylor, P. C. *Tetrahedron Lett.* **1998**, 39, 8509. (c) Jonas, R. T.; Stack, T. D. P. *Inorg. Chem.* **1998**, 37, 6615.
- (61) Ishimori, K.; Watanabe, M.; Kimura, T.; Murata, M.; Nishihara, H. *J. Alloys Compd.* **2006**, 408–412, 1278.
- (62) Watanabe, M.; Nankawa, T.; Yamada, T.; Kimura, T.; Namiki, K.; Murata, M.; Nishihara, H.; Tachimori, S. *Inorg. Chem.* **2003**, 42, 6977.
- (63) Case, F. H. *J. Het. Chem.* **1971**, 8, 1043.
- (64) (a) Pabst, G. R.; Sauer, J. *Tetrahedron Lett.* **1998**, 39, 6687. (b) Sauer, J.; Heldmann, D. K.; Pabst, G. R. *Eur. J. Org. Chem.* **1999**, 313. (c) Pabst, G. R.; Pfüller, O. C.; Sauer, J. *Tetrahedron* **1999**, 55, 5047. (d) Pabst, G. R.; Sauer, J. *Tetrahedron* **1999**, 55, 5067.
- (65) Kolarik, Z.; Müllich, U.; Gassner, F. *Solvent Extr. Ion Exch.* **1999**, 17, 23.
- (66) Weigl, M.; Müllich, U.; Geist, A.; Gompfer, K.; Zevaco, T.; Stephan, H. *J. Rad. Nucl. Chem.* **2003**, 256, 403.
- (67) Kolarik, Z.; Müllich, U.; Gassner, F. *Solvent Extr. Ion Exch.* **1999**, 17, 1155.
- (68) Weigl, M.; Geist, A.; Müllich, U.; Gompfer, K. *Solvent Extr. Ion Exch.* **2006**, 24, 845.
- (69) Drew, M. G. B.; Guillauneux, D.; Hudson, M. J.; Iveson, P. B.; Madic, C. *Inorg. Chem. Commun.* **2001**, 4, 462.
- (70) Boucher, C.; Drew, M. G. B.; Giddings, P.; Harwood, L. M.; Hudson, M. J.; Iveson, P. B.; Madic, C. *Inorg. Chem. Commun.* **2002**, 5, 596.
- (71) Drew, M. G. B.; Hudson, M. J.; Iveson, P. B.; Madic, C.; Russell, M. L. *J. Chem. Soc. Dalton Trans.* **1999**, 2433.

- (72) Drew, M. G. B.; Guillaneux, D.; Hudson, M. J.; Iveson, P. B.; Russell, M. L.; Madic, C. *Inorg. Chem. Commun.* **2001**, 4, 12.
- (73) (a) Iveson, P. B.; Rivière, C.; Guillaneux, D.; Nierlich, M.; Thuéry, P.; Ephritikhine, M.; Madic, C. *Chem. Commun.* **2001**, 1512. (b) Berthet, J.-C.; Miquel, Y.; Iveson, P. B.; Nierlich, M.; Thuéry, P.; Madic, C.; Ephritikhine, M. *J. Chem. Soc. Dalton Trans.* **2002**, 3265.
- (74) Benay, G.; Schurhammer, R.; Desaphy, J.; Wipff, G. *New J. Chem.* **2011**, 35, 184.
- (75) (a) Colette, S.; Amekraz, B.; Madic, C.; Berthon, L.; Cote, G.; Moulin, C. *Inorg. Chem.* **2002**, 41, 7031. (b) Colette, S.; Amekraz, B.; Madic, C.; Berthon, L.; Cote, G.; Moulin, C. *Inorg. Chem.* **2003**, 42, 2215. (c) Colette, S.; Amekraz, B.; Madic, C.; Berthon, L.; Cote, G.; Moulin, C. *Inorg. Chem.* **2004**, 43, 6745. (d) Steppert, M.; Walther, C.; Geist, A.; Fanghänel, T. *New J. Chem.* **2009**, 33, 2437.
- (76) (a) Denecke, M. A.; Rossberg, A.; Panak, P. J.; Weigl, M.; Schimmelpfennig, B.; Geist, A. *Inorg. Chem.* **2005**, 44, 8418. (b) Miguiditchian, M.; Guillaneux, D.; Francois, N.; Airvault, S.; Ducros, S.; Thauvin, D.; Madic, C.; Illemassène, M.; Lagarde, G.; Krupa, J.-C. *Nucl. Sci. Eng.* **2006**, 153, 223. (c) Trumm, S.; Panak, P. J.; Geist, A.; Fanghänel, T. *Eur. J. Inorg. Chem.* **2010**, 3022. (d) Banik, N. L.; Schimmelpfennig, B.; Marquardt, C. M.; Brendebach, B.; Geist, A.; Denecke, M. A. *Dalton Trans.* **2010**, 39, 5117.
- (77) Berthet, J.-C.; Thuéry, P.; Dognon, J.-P.; Guillaneux, D.; Ephritikhine, M. *Inorg. Chem.* **2008**, 47, 6850.
- (78) Ionova, G.; Rabbe, C.; Guillaumont, R.; Ionov, S.; Madic, C.; Krupa, J.-C.; Guillaneux, D. *New J. Chem.* **2002**, 26, 234.
- (79) Benay, G.; Schurhammer, R.; Wipff, G. *Phys. Chem. Chem. Phys.* **2010**, 12, 11089.
- (80) (a) Chevrot, G.; Schurhammer, R.; Wipff, G. *Phys. Chem. Chem. Phys.* **2007**, 9, 1991. (b) Chevrot, G.; Schurhammer, R.; Wipff, G. *Phys. Chem. Chem. Phys.* **2007**, 9, 5928.
- (81) (a) Madic, C.; Hudson, M. J.; Liljenzin, J.-O.; Glatz, J.-P.; Nannicini, R.; Facchini, A.; Kolarik, Z.; Odoj, R. *Prog. Nucl. Energ.* **2002**, 40, 523. (b) Madic, C.; Boullis, B.; Baron, P.; Testard, F.; Hudson, M. J.; Liljenzin, J.-O.; Christiansen, B.; Ferrando, M.; Facchini, A.; Geist, A.; Modolo, G.; Espartero, A. G.; De Mendoza, J. J. *Alloys Compd.* **2007**, 444–445, 23.
- (82) (a) Wei, Y.-Z.; Hoshi, H.; Kumagai, M.; Asakura, T.; Morita, Y. *J. Alloys Compd.* **2004**, 374, 447. (b) Bhattacharyya, A.; Mohapatra, P. K.; Roy, A.; Gadly, T.; Ghosh, S. K.; Manchanda, V. K. *Hydrometallurgy* **2009**, 99, 18.
- (83) (a) Madic, C.; Hudson, M. J.; Liljenzin, J.-O.; Glatz, J.-P.; Nannicini, R.; Facchini, A.; Kolarik, Z.; Odoj, R. *New Partitioning Techniques for Minor Actinides*. EUR19149EN (2000). (b) Hill, C.; Guillaneux, D.; Berthon, L.; Madic, C. *J. Nucl. Sci. Technol.* **2002**, Supplement 3, 309.
- (84) Nilsson, M.; Andersson, S.; Ekberg, C.; Foreman, M. R. S.; Hudson, M. J.; Skarnemark, G. *Radiachim. Acta* **2006**, 94, 103.
- (85) Hudson, M. J.; Foreman, M. R. St. J.; Hill, C.; Huet, N.; Madic, C. *Solvent Extr. Ion Exch.* **2003**, 21, 637.
- (86) Case, F. H. *J. Org. Chem.* **1965**, 30, 931.
- (87) (a) Fife, W. K. *J. Org. Chem.* **1983**, 48, 1375. (b) Norrby, T.; Börje, A.; Zhang, L.; Åkermark, B. *Acta Chem. Scand.* **1998**, 52, 77.
- (88) Hudson, M. J.; Drew, M. G. B.; Foreman, M. R. St. J.; Hill, C.; Huet, N.; Madic, C.; Youngs, T. G. A. *Dalton Trans.* **2003**, 1675.
- (89) Krejzler, J.; Narbutt, J.; Foreman, M. R. St. J.; Hudson, M. J.; Casensky, B.; Madic, C. *Czech. J. Phys.* **2006**, 56, D459.
- (90) Girnt, D.; Roesky, P. W.; Geist, A.; Ruff, C. M.; Panak, P. J.; Denecke, M. A. *Inorg. Chem.* **2010**, 49, 9627.
- (91) Drew, M. G. B.; Foreman, M. R. St. J.; Geist, A.; Hudson, M. J.; Marken, F.; Norman, V.; Weigl, M. *Polyhedron* **2006**, 25, 888.
- (92) Drew, M. G. B.; Hill, C.; Hudson, M. J.; Iveson, P. B.; Madic, C.; Youngs, T. G. A. *Dalton Trans.* **2004**, 244.
- (93) p
- (94) p
- (95) Drew, M. G. B.; Foreman, M. R. S. J.; Hill, C.; Hudson, M. J.; Madic, C. *Inorg. Chem. Commun.* **2005**, 8, 239.
- (96) Narbutt, J.; Krejzler, J. *Radiachim. Acta* **2008**, 96, 219.
- (97) Foreman, M. R. S.; Hudson, M. J.; Drew, M. G. B.; Hill, C.; Madic, C. *Dalton Trans.* **2006**, 1645.
- (98) Hubscher-Bruder, V.; Haddaoui, J.; Bouhroum, S.; Arnaud-Neu, F. *Inorg. Chem.* **2010**, 49, 1363.
- (99) Retegan, T.; Berthon, L.; Ekberg, C.; Fermvik, A.; Skarnemark, G.; Zorz, N. *Solvent Extr. Ion Exch.* **2009**, 27, 663.
- (100) (a) Foreman, M. R. St. J.; Hudson, M. J.; Geist, A.; Madic, C.; Weigl, M. *Solvent Extr. Ion Exch.* **2005**, 23, 645. (b) Nilsson, M.; Andersson, S.; Drouet, F.; Ekberg, C.; Foreman, M.; Hudson, M.; Liljenzin, J.-O.; Magnusson, D.; Skarnemark, G. *Solvent Extr. Ion Exch.* **2006**, 24, 299.
- (101) Nilsson, M.; Ekberg, C.; Foreman, M.; Hudson, M.; Liljenzin, J.-O.; Modolo, G.; Skarnemark, G. *Solvent Extr. Ion Exch.* **2006**, 24, 823.
- (102) Retegan, T.; Ekberg, C.; Dubois, I.; Fermvik, A.; Skarnemark, G.; Wass, T. J. *Solvent Extr. Ion Exch.* **2007**, 25, 417.
- (103) Ekberg, C.; Dubois, I.; Fermvik, A.; Retegan, T.; Skarnemark, G.; Drew, M. G. B.; Foreman, M. R. S.; Hudson, M. J. *Solvent Extr. Ion Exch.* **2007**, 25, 603.
- (104) Geist, A.; Hill, C.; Modolo, G.; Foreman, M. R. St. J.; Weigl, M.; Gompfer, K.; Hudson, M. J. *Solvent Extr. Ion Exch.* **2006**, 24, 463.
- (105) Magnusson, D.; Christiansen, B.; Glatz, J.-P.; Malmbeck, R.; Modolo, G.; Serrano-Purroy, D.; Sorel, C. *Radiachim. Acta* **2009**, 97, 155.
- (106) (a) Malmbeck, R.; Magnusson, D.; Glatz, J.-P. in *Proceedings of Global 2009*, Paris, France, **2009**, paper 1937. (b) Magnusson, D.; Christiansen, B.; Foreman, M. R. S.; Geist, A.; Glatz, J.-P.; Malmbeck, R.; Modolo, G.; Serrano-Purroy, D.; Sorel, C. *Solvent Extr. Ion Exch.* **2009**, 27, 97.
- (107) Aneheim, E.; Ekberg, C.; Fermvik, A.; Foreman, M. R. St. J.; Retegan, T.; Skarnemark, G. *Solvent Extr. Ion Exch.* **2010**, 28, 437.
- (108) Berthet, J.-C.; Thuéry, P.; Foreman, M. R. S.; Ephritikhine, M. *Radiachim. Acta* **2008**, 96, 189.
- (109) (a) Berthet, J.-C.; Siffredi, G.; Thuéry, P.; Ephritikhine, M. *Dalton Trans.* **2009**, 3478. (b) Berthet, J.-C.; Maynadié, J.; Thuéry, P.; Ephritikhine, M. *Dalton Trans.* **2010**, 39, 6801.
- (110) Trumm, S.; Lieser, G.; Foreman, M. R. S. J.; Panak, P. J.; Geist, A.; Fanghänel, T. *Dalton Trans.* **2010**, 39, 923.
- (111) Benay, G.; Schurhammer, R.; Wipff, G. *Phys. Chem. Chem. Phys.* **2011**, 13, 2922.
- (112) Lewis, F. W.; Harwood, L. M.; Hudson, M. J.; Drew, M. G. B.; Modolo, G.; Sypula, M.; Desreux, J. F.; Bouslimani, N.; Vidick, G. *Dalton Trans.* **2010**, 39, 5172.
- (113) (a) Mincher, B. J.; Modolo, G.; Mezyk, S. P. *Solvent Extr. Ion Exch.* **2009**, 27, 1. (b) Mincher, B. J.; Modolo,

- G.; Mezyk, S. P. *Solvent Extr. Ion Exch.* **2009**, 27, 331.
(c) Mincher, B. J.; Modolo, G.; Mezyk, S. P. *Solvent Extr. Ion Exch.* **2009**, 27, 579. (d) Mincher, B. J.; Modolo, G.; Mezyk, S. P. *Solvent Extr. Ion Exch.* **2010**, 28, 415.
- (114) Buxton, G. V.; Greenstock, C. L.; Helman, W. P.; Ross, A. B. *J. Phys. Chem. Ref. Data* **1988**, 17, 513.
- (115) Spinks, J. W. T.; Woods, R. J. in *An Introduction to Radiation Chemistry*, 3rd ed, John Wiley and Sons: New York, **1990**.
- (116) Fermvik, A.; Berthon, L.; Ekberg, C.; Englund, S.; Retegan, T.; Zorz, N. *Dalton Trans.* **2009**, 6421.
- (117) Hudson, M. J.; Boucher, C. E.; Braekers, D.; Desreux, J. F.; Drew, M. G. B.; Foreman, M. R. St J.; Harwood, L. M.; Hill, C.; Madic, C.; Marken, F.; Youngs, T. G. A. *New J. Chem.* **2006**, 30, 1171.
- (118) Ekberg, C.; Aneheim, E.; Fermvik, A.; Foreman, M.; Löfström-Engdahl, E.; Retegan, T.; Spendlikova, I. *J. Chem. Eng. Data* **2010**, 55, 5133.
- (119) Retegan, T.; Ekberg, C.; Englund, S.; Fermvik, A.; Foreman, M. R. S.; Skarnemark, G. *Radiochim. Acta* **2007**, 95, 637.
- (120) Fermvik, A.; Ekberg, C.; Englund, S.; Foreman, M. R. S. J.; Modolo, G.; Retegan, T.; Skarnemark, G. *Radiochim. Acta* **2009**, 97, 319.
- (121) Magnusson, D.; Christiansen, B.; Malmbeck, R.; Glatz, J.-P. *Radiochim. Acta* **2009**, 97, 497.
- (122) (a) Coffman, D. D.; Jenner, E. L.; Lipscomb, R. D. *J. Am. Chem. Soc.* **1958**, 80, 2864. (b) Bremner, D. H.; Mitchell, S. R.; Staines, H. *Ultrasonics Sonochemistry* **1996**, 3, 47. (c) Bremner, D. H.; Burgess, A. E.; Li, F.-B. *Green Chem.* **2001**, 3, 126.
- (123) (a) Jones, P.; Villeneuve, G. B.; Fei, C.; DeMarte, J.; Haggarty, A. J.; Nwe, K. T.; Martin, D. A.; Lebuis, A.-M.; Finkelstein, J. M.; Gour-Salin, B. J.; Chan, T. H.; Leyland-Jones, B. R. *J. Med. Chem.* **1998**, 41, 3062. (b) Kikuchi, K.; Hibi, S.; Yoshimura, H.; Tokuhara, N.; Tai, K.; Hida, T.; Yamauchi, T.; Nagai, M. *J. Med. Chem.* **2000**, 43, 409.
- (124) Harwood, L. M.; Hudson, M. J.; Lewis, F. W. GB0922588.9 (**2009**).
- (125) For a similar reaction, see: Ibarzo, J.; Ortuña, R. M. *Tetrahedron* **1994**, 50, 9825.

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Laurence M. Harwood ppp

Development of Highly Selective Ligands for Separations of Actinides from Lanthanides in the Nuclear Fuel Cycle

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² Ref. 3 in 464

³ Ref. 2 in our JACS paper, Ref. 6 in 464

⁴ Ref. 3 in our JACS paper

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⁷ Ref. 2 in A, Ref. 10 in 380

⁸ Ref. 7 in 466

⁹ Ref. 3 in A, Ref. 66 in 466

¹⁰ Ref. 9 in 380

¹¹ Ref. 26 in our JACS paper

¹² Ref. 1 in A, Ref. 77 in 466

¹³ Ref. 5 in our JACS paper

¹⁴ Ref. 6 in A

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¹²⁵ For a similar reaction, see Ref. 20 in our JACS paper